

CHEMICAL BONDS

IONIC, COVALENT, & METALLIC

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READING COMPREHENSION & QUESTIONS

SCIENCE REAL

Types of Chemical Bonds Reading Comprehension

What is the difference between a molecule and a compound?

What of the substances shown in the diagram above are molecules?

What of the substances shown in the diagram above are compounds?

Chemical Bonds and Electronegativity

Electrons in atoms exist in two main energy levels, or shells. Electrons closest to the nucleus are most strongly attracted to the nucleus and are called inner electrons. Electrons farthest from the nucleus are called outer electrons, and they tend to be more easily lost or shared with other atoms. Typically, atoms gain or lose electrons to form ionic bonds or share electrons to form covalent bonds.

The two most common types of chemical bonds are ionic bonds and covalent bonds. Ionic bonds form when one atom gives up an electron to another atom, while covalent bonds form when two atoms share electrons.

Whether a bond is ionic or covalent typically depends on the electronegativity of the atoms involved. Electronegativity is an atom's ability to attract electrons. Elements with low electronegativity have outer shells that are almost empty and tend to gain or lose electrons easily. Elements with high electronegativity have outer shells that are nearly full and tend to hold on to their electrons. Ionic bonds form when the electronegativity of one atom is very high, and the electronegativity of the other atom is very low. Covalent bonds form when the electronegativity of the atoms are more similar. Typically, atoms with high electronegativity have outer shells that are nearly full, and atoms with low electronegativity have outer shells that are nearly empty.

Electronegativity increases from bottom left to top right on the periodic table.

Diagram A: Oxygen atom has a higher electronegativity than Sodium. Which one is a cation and which one is an anion?

Diagram B: Which one is a cation and which one is an anion?

Diagram C: Which one is a cation and which one is an anion?

Diagram D: Which one is a cation and which one is an anion?

Diagram E: Which one is a cation and which one is an anion?

Diagram F: Which one is a cation and which one is an anion?

Diagram G: Which one is a cation and which one is an anion?

Diagram H: Which one is a cation and which one is an anion?

Diagram I: Which one is a cation and which one is an anion?

Diagram J: Which one is a cation and which one is an anion?

Diagram K: Which one is a cation and which one is an anion?

Diagram L: Which one is a cation and which one is an anion?

Diagram M: Which one is a cation and which one is an anion?

Diagram N: Which one is a cation and which one is an anion?

Diagram O: Which one is a cation and which one is an anion?

Diagram P: Which one is a cation and which one is an anion?

Diagram Q: Which one is a cation and which one is an anion?

Diagram R: Which one is a cation and which one is an anion?

Diagram S: Which one is a cation and which one is an anion?

Diagram T: Which one is a cation and which one is an anion?

Diagram U: Which one is a cation and which one is an anion?

Diagram V: Which one is a cation and which one is an anion?

Diagram W: Which one is a cation and which one is an anion?

Diagram X: Which one is a cation and which one is an anion?

Diagram Y: Which one is a cation and which one is an anion?

Diagram Z: Which one is a cation and which one is an anion?

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Bonding In Metals Guided Study Answers

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