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web a write a balanced equation for the reaction al o<sub>2</sub> al<sub>2</sub>o<sub>3</sub> 4 al 3 o<sub>2</sub> 2 al<sub>2</sub>o<sub>3</sub> b determine the theoretical yield of al<sub>2</sub>o<sub>3</sub> 160 0 g al x 1 mol al x 2 mol al<sub>2</sub>o<sub>3</sub> x 101 964 g al<sub>2</sub>o<sub>3</sub> 302 3 g al<sub>2</sub>o<sub>3</sub> 26 982 g al 4 mol al 1 mol al<sub>2</sub>o<sub>3</sub> c determine the percent yield yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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