

**CHEM 1412 Practice exam 1 Chapter 13-15**

1. Which of the following concentration unit will be affected by temperature?  
 (A) mass percent (B) **molarity** (C) mole fraction (D) molality

2. Which of the following action will reduce the solubility of a gas in water?  
 (A) pressure (B) molar mass of the gas (C) acidic or basic gas (D) **temperature**

**Answer questions Q3-Q6 based on the following information**

A solution of 12.00 g of an unknown nondissociating compound dissolved in 200.0 g of benzene freezes at 3.45°C. The  $K_f$  of benzene is 5.07°C/m. The freezing point of benzene is 5.45°C.

3. What is the mass percent of the solution?

$$12.00\text{g}/(12.00+200.0)\text{g}=5.66\%$$

4. What is the molality of the unknown solution?

$$\Delta T_f = K_f m, \Delta T_f = T_f^\circ - T_f = 5.45 - 3.45 = 2.00^\circ\text{C}, m = \Delta T_f / K_f = 2.00^\circ\text{C} / (5.07^\circ\text{C}/m) = 0.394m$$

5. How many moles of unknown in the solution?

The molality  $m$  = moles of solute/1kg solvent, in 200.0g of solvent, there are 200.0g x (1kg/1000g) x 0.394mol/ kg solvent = 0.0789 mol

6. What is the molar mass of the unknown?

$$M = \text{mass}/n = 12.00\text{g}/0.0789 \text{ mol} = 152.09 \text{ g/mol}$$

7. The osmotic pressure of 0.050 M  $\text{MgSO}_4$  at 25 °C is 1.60 atm. What is the Van't Hoff factor?

(A) 1.9 (B) **1.3** (C) 2.5 (D) 2

$$\pi = iMRT, i = \pi / MRT = 1.60 \text{ atm} / [0.050\text{M} \times (0.082\text{atm} \cdot \text{L}/\text{K} \cdot \text{mol}) \times 298\text{K}] = 1.3$$

8. The freezing point of 21.2 g of NaCl in 135 mL of water ( $K_f$  = 1.86°C/m) is \_\_\_\_\_.

(A) -5.00 °C (B) -4.60 °C (C) **-9.99 °C** (D) -11.5 °C

9. The rate at which hydrogen reacts in the reaction  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$  is  $0.074 \text{ Ms}^{-1}$ . The rate of formation of ammonia is \_\_\_\_\_  $\text{Ms}^{-1}$ .

A) 0.148

B) 0.049

C) 0.111

D) 0.222

**Strategy:** The rate is defined as the change in concentration of a reactant or product with time. Each "change in concentration" term is divided by the corresponding stoichiometric coefficient. Terms involving reactants are preceded by a minus sign.

$$\text{rate} = -\frac{\Delta[\text{N}_2]}{\Delta t} = -\frac{1}{3} \frac{\Delta[\text{H}_2]}{\Delta t} = \frac{1}{2} \frac{\Delta[\text{NH}_3]}{\Delta t}$$

**Solution:**

(a) If hydrogen is reacting at the rate of  $-0.074 \text{ Ms}^{-1}$ , the rate at which ammonia is being formed is

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