

More aqueous equilibrium self test/practice problems, Dr. Scott Ensign.

1. **(Acid-base equilibrium review)** A solution is prepared by dissolving 0.15 moles of acetic acid in 1.00 liters of solution. K_a for acetic acid is 1.8×10^{-5} . What will be the pH of the solution once equilibrium is established?

- a. 2.0
- b. 5.6
- c. 2.8
- d. 3.9
- e. 4.7

2. **(Acid-base equilibrium review)** A solution is prepared by dissolving 0.10 moles of sodium acetate (NaCH_3COO) in 1.00 liters of solution. K_a for acetic acid is 1.8×10^{-5} . What will be the pH of the solution once equilibrium is established?

- a. 8.9
- b. 5.1
- c. 9.9
- d. 4.1
- e. 9.3

3. A solution is prepared by mixing 0.15 moles of acetic acid (CH_3COOH) with 0.10 moles of sodium acetate (NaCH_3COO) in 1.00 liters of solution. K_a for acetic acid is 1.8×10^{-5} . What will be the pH of the solution once equilibrium is established?

- a. 4.9
- b. 6.6
- c. 2.9
- d. 4.6
- e. 4.7

4. A solution is prepared by mixing 0.15 moles of acetic acid (CH_3COOH) with 0.050 moles of nitric acid (HNO_3) in 1.00 liters of solution. K_a for acetic acid is 1.8×10^{-5} . What will be the pH of the solution once equilibrium is established?

- a. 2.4
- b. 2.8
- c. 1.3
- d. 4.3
- e. 5.2

5. **(Acid-base equilibrium review)** What would be the pH of a solution prepared by dissolving 0.55 moles of NH_3 in 1.00 liters of solution? K_b for NH_3 is 1.80×10^{-5}

- a. 9.51
- b. 11.5
- c. 11.8
- d. 2.50
- e. 9.00

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Stephen DeMeo



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