

- According to the Bronsted—Lowry definition, which chemical species can function *both* as an acid and as a base?
  - $\text{Cl}^-$
  - $\text{SO}_4^{2-}$
  - $\text{NH}_4^+$
  - $\text{HCO}_3^-$
  - $\text{H}_3\text{O}^+$
- In the reaction  $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$  which is an acid—base conjugate pair?
  - $\text{H}_2\text{O}$  and  $\text{HCN}$
  - $\text{H}_2\text{O}$  and  $\text{OH}^-$
  - $\text{CN}^-$  and  $\text{H}_2\text{O}$
  - $\text{HCN}$  and  $\text{OH}^-$
- Given that  $\text{HX}$  is a stronger Bronsted acid than  $\text{HY}$  in aqueous solution, which is true of a 1 M solution of  $\text{NaX}$ ?
  - It is less basic than a 1 M solution of  $\text{NaY}$ .
  - It is more basic than a 1 M solution of  $\text{NaY}$ .
  - It yields a neutral solution.
  - It is more concentrated than a 1 M solution of  $\text{NaY}$ .
- The pH of a 0.03 M  $\text{HCl}$  solution is
  - 1.5
  - 2.5
  - 3.5
  - 12.5
- The pH of a  $1.0 \times 10^{-2}$  M  $\text{Ba}(\text{OH})_2$  solution at  $25^\circ\text{C}$  is
  - 2.7
  - 3.0
  - 11.0
  - 11.3
- Which is the strongest acid?
  - $\text{HClO}$
  - $\text{HClO}_2$
  - $\text{HClO}_3$
  - $\text{HClO}_4$
- Which salt reacts with water (hydrolyzes) to produce a basic solution?
  - $\text{NaC}_2\text{H}_3\text{O}_2$
  - $\text{NaNO}_3$
  - $\text{NH}_4\text{Cl}$
  - $\text{BaSO}_4$

- The *weakest* of the bases listed is
 

| Acid                 | Conjugate Base     | $K_a$ (Ionization Constant of Acid) |
|----------------------|--------------------|-------------------------------------|
| $\text{HCl}$         | $\text{Cl}^-$      | 100% ionized                        |
| $\text{HSO}_4^-$     | $\text{SO}_4^{2-}$ | $1.2 \times 10^{-2}$                |
| $\text{H}_2\text{S}$ | $\text{HS}^-$      | $5.7 \times 10^{-8}$                |
| $\text{HS}^-$        | $\text{S}^{2-}$    | $1.2 \times 10^{-13}$               |

  - $\text{Cl}^-$
  - $\text{CN}^-$
  - $\text{HS}^-$
  - $\text{S}^{2-}$
  - $\text{SO}_4^{2-}$

- The oxide of which element will react with water to form the strongest acid?
 

|        | Main Groups |    |     |    |   |    |     |   |
|--------|-------------|----|-----|----|---|----|-----|---|
| Period | I           | II | III | IV | V | VI | VII | O |
| First  |             |    |     |    |   |    |     |   |
| Second | X           | Y  | Z   | P  |   | Q  | S   | U |
| Third  | W           |    |     |    |   | R  | T   | M |

  - W
  - M
  - P
  - R
  - Z

- Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate  $\text{KC}_2\text{H}_3\text{O}_2$  is less alkaline than a 0.10 M solution of potassium cyanide,  $\text{KCN}$ ?
  - Hydrocyanic acid is a weaker acid than acetic acid.
  - Hydrocyanic acid is less soluble in water than acetic acid.
  - Cyanides are less soluble than acetates.
  - Acetic acid is a weaker acid than hydrocyanic acid.
- In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M  $\text{NaOH}$ , the properties of the solution at the equivalence point will correspond exactly to the properties of
  - a 0.100 M sodium solution.
  - a 0.0500 M sodium hydroxide solution.
  - a 0.0500 M benzoic acid solution.
  - a 0.0500 M sodium benzoate solution.

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## **Table of Contents Acs Quantitative Analysis Practice Exam**

1. Understanding the eBook Acs Quantitative Analysis Practice Exam
  - The Rise of Digital Reading Acs Quantitative Analysis Practice Exam
  - Advantages of eBooks Over Traditional Books
2. Identifying Acs Quantitative Analysis Practice Exam
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acs Quantitative Analysis Practice Exam
  - User-Friendly Interface
4. Exploring eBook Recommendations from Acs Quantitative Analysis Practice Exam
  - Personalized Recommendations
  - Acs Quantitative Analysis Practice Exam User Reviews and Ratings
  - Acs Quantitative Analysis Practice Exam and Bestseller Lists
5. Accessing Acs Quantitative Analysis Practice Exam Free and Paid eBooks
  - Acs Quantitative Analysis Practice Exam Public Domain eBooks
  - Acs Quantitative Analysis Practice Exam eBook Subscription Services
  - Acs Quantitative Analysis Practice Exam Budget-Friendly Options
6. Navigating Acs Quantitative Analysis Practice Exam eBook Formats

- ePub, PDF, MOBI, and More
- Acs Quantitative Analysis Practice Exam Compatibility with Devices
- Acs Quantitative Analysis Practice Exam Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Acs Quantitative Analysis Practice Exam
  - Highlighting and Note-Taking Acs Quantitative Analysis Practice Exam
  - Interactive Elements Acs Quantitative Analysis Practice Exam
- 8. Staying Engaged with Acs Quantitative Analysis Practice Exam
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Acs Quantitative Analysis Practice Exam
- 9. Balancing eBooks and Physical Books Acs Quantitative Analysis Practice Exam
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Acs Quantitative Analysis Practice Exam
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Acs Quantitative Analysis Practice Exam
  - Setting Reading Goals Acs Quantitative Analysis Practice Exam
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acs Quantitative Analysis Practice Exam
  - Fact-Checking eBook Content of Acs Quantitative Analysis Practice Exam
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
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