

Name \_\_\_\_\_ Section \_\_\_\_\_

## Acid-Base Titration (Volumetric Analysis) – Data Sheet

CHEM 161A – General Chemistry II Lab

Attach all calculations as separate pages.

### Part C – Standardization of NaOH

$$M_a V_a = M_b V_b$$

Standardization of NaOH	Trial 1	Trial 2
Mass $\text{KC}_8\text{H}_5\text{O}_4$ (KHP) weighed (g)	0.4275(g)	0.4275(g)
Molar mass $\text{KC}_8\text{H}_5\text{O}_4$ (KHP) (g/mol)	204.22121(g/mol)	204.22121(g/mol)
Moles $\text{KC}_8\text{H}_5\text{O}_4$ (KHP) used (mol) <span style="margin-left: 20px;">(MaVa)</span>	0.002093 mol	0.002093 mol
Stoichiometric ratio of KHP to NaOH	1 mol KHP:1 mol NaOH	1 mol KHP:1 mol NaOH
Moles NaOH neutralized (mol)	+	
Initial volume reading NaOH from buret (mL)	20.01 mL	20.01 mL
Final volume reading NaOH from buret (mL)	37.26 mL	37.26 mL
Volume NaOH used (mL) <span style="margin-left: 20px;">(V)</span>	17.25 mL	17.25 mL
Molarity of NaOH (mol/L)		
Average molarity of NaOH (mol/L)		

= ?  
2

### Part D – Titration of an Unknown Diprotic Acid

Molar Mass of an Unknown Acid	Trial 1	Trial 2
Mass of unknown acid (g)	0.1072 g	0.1244 g
Standardized molarity of NaOH (mol/L) [Part C]		
Initial volume reading NaOH from buret (mL)	0.31 mL	0.25 mL
Final volume reading NaOH from buret (mL)	18.53 mL	21.72 mL
Volume NaOH used (mL)		
Moles NaOH used (mol)		
Stoichiometric ratio of NaOH to acid		
Moles of acid neutralized (mol)		
Molar mass unknown acid (g/mol)		
Average molar mass unknown acid (g/mol)		

(Answer)

# Acid Base Titration And Volumetric Analysis Lab 42

## Answers

**Arthur James Wells**



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## **Table of Contents Acid Base Titration And Volumetric Analysis Lab 42 Answers**

1. Understanding the eBook Acid Base Titration And Volumetric Analysis Lab 42 Answers
  - The Rise of Digital Reading Acid Base Titration And Volumetric Analysis Lab 42 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Titration And Volumetric Analysis Lab 42 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acid Base Titration And Volumetric Analysis Lab 42 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Base Titration And Volumetric Analysis Lab 42 Answers
  - Personalized Recommendations
  - Acid Base Titration And Volumetric Analysis Lab 42 Answers User Reviews and Ratings
  - Acid Base Titration And Volumetric Analysis Lab 42 Answers and Bestseller Lists
5. Accessing Acid Base Titration And Volumetric Analysis Lab 42 Answers Free and Paid eBooks
  - Acid Base Titration And Volumetric Analysis Lab 42 Answers Public Domain eBooks
  - Acid Base Titration And Volumetric Analysis Lab 42 Answers eBook Subscription Services
  - Acid Base Titration And Volumetric Analysis Lab 42 Answers Budget-Friendly Options
6. Navigating Acid Base Titration And Volumetric Analysis Lab 42 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Acid Base Titration And Volumetric Analysis Lab 42 Answers Compatibility with Devices
  - Acid Base Titration And Volumetric Analysis Lab 42 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Acid Base Titration And Volumetric Analysis Lab 42 Answers
  - Highlighting and Note-Taking Acid Base Titration And Volumetric Analysis Lab 42 Answers
  - Interactive Elements Acid Base Titration And Volumetric Analysis Lab 42 Answers
8. Staying Engaged with Acid Base Titration And Volumetric Analysis Lab 42 Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Acid Base Titration And Volumetric Analysis Lab 42 Answers
  9. Balancing eBooks and Physical Books Acid Base Titration And Volumetric Analysis Lab 42 Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Acid Base Titration And Volumetric Analysis Lab 42 Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Acid Base Titration And Volumetric Analysis Lab 42 Answers
    - Setting Reading Goals Acid Base Titration And Volumetric Analysis Lab 42 Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Acid Base Titration And Volumetric Analysis Lab 42 Answers
    - Fact-Checking eBook Content of Acid Base Titration And Volumetric Analysis Lab 42 Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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