

11. Which statement about one atom of an element identifies the element?
- A) The atom has 1 proton.
B) The atom has 2 neutrons.
C) The sum of the number of protons and neutrons in the atom is 3.
D) The difference between the number of neutrons and protons in the atom is 1.
12. Compared to a proton, an electron has
- A) a greater quantity of charge and the same sign
B) a greater quantity of charge and the opposite sign
C) the same quantity of charge and the same sign
D) the same quantity of charge and the opposite sign
13. Every chlorine atom has
- A) 7 electrons
B) 17 neutrons
C) a mass number of 35
D) an atomic number of 17
14. What can be determined if only the atomic number of an atom is known?
- A) the total number of neutrons in the atom, only
B) the total number of protons in the atom, only
C) the total number of protons and the total number of neutrons in the atom
D) the total number of protons and the total number of electrons in the atom
15. Which of the following atoms has the greatest nuclear charge?
- A) ${}^{14}_7\text{N}$ B) ${}^{12}_6\text{C}$ C) ${}^7_3\text{H}$ D) ${}^4_2\text{He}$
16. What is the total number of neutrons in the nucleus of a neutral atom that has 19 electrons and a mass number of 39?
- A) 19 B) 20 C) 39 D) 58
17. The weighted average of the atomic masses of the naturally occurring isotopes of an element is the
- A) atomic mass of the element
B) atomic number of the element
C) mass number of each isotope
D) formula mass of each isotope
18. What is the mass number of an atom that consists of 20 protons, 20 neutrons, and 18 electrons?
- A) 18 B) 20 C) 38 D) 40
19. An atom that contains 8 protons, 8 electrons, and 9 neutrons has
- A) an atomic number of 9
B) an atomic number of 16
C) a mass number of 17
D) a mass number of 25
20. Compared to an atom of C-12, an atom of C-14 has a greater
- A) number of electrons
B) number of protons
C) atomic number
D) mass number
21. Atoms of different isotopes of the same element differ in their total number of
- A) electrons B) neutrons
C) protons D) valence electrons
22. Which particle has two neutrons?
- A) ${}^1_0\text{n}$ B) ${}^1_1\text{H}$ C) ${}^2_1\text{H}$ D) ${}^4_2\text{He}$

Atomic Structure Test Answers

Xiang Xie



Atomic Structure Test Answers:

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Table of Contents Atomic Structure Test Answers

1. Understanding the eBook Atomic Structure Test Answers
 - The Rise of Digital Reading Atomic Structure Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Atomic Structure Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Atomic Structure Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Atomic Structure Test Answers
 - Personalized Recommendations
 - Atomic Structure Test Answers User Reviews and Ratings
 - Atomic Structure Test Answers and Bestseller Lists
5. Accessing Atomic Structure Test Answers Free and Paid eBooks
 - Atomic Structure Test Answers Public Domain eBooks
 - Atomic Structure Test Answers eBook Subscription Services
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 - Interactive Elements Atomic Structure Test Answers
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Atomic Structure Test Answers
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Carving Out Dedicated Reading Time
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 - Utilizing eBooks for Skill Development
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