

## Topic Name : Atomic Structure

1. The mass of a proton is equal to \_\_\_\_\_.  
 (1) 0.0055 u (2) 1.00727 u  
 (3) 1.6000 u (4) 1.672 u
2. Cathode rays are made up of \_\_\_\_\_.  
 (1) positively charged particles  
 (2) negatively charged particles  
 (3) neutral particles  
 (4) equal amount of positively and negatively charged particles
3. An  $\alpha$ -particle is represented as \_\_\_\_\_.  
 (1)  ${}_1^4\alpha$  (2)  ${}_2^4\alpha$   
 (3)  ${}_2^4\alpha$  (4)  ${}_2^4\alpha$
4. The mass number and atomic number of an element are A and Z respectively. The number of neutrons in the atom of that element is calculated as \_\_\_\_\_.  
 (1) A (2) Z  
 (3) A + Z (4) A - Z
5. The nuclide having 6 protons and 8 neutrons is represented as \_\_\_\_\_.  
 (1)  ${}_6^{14}\text{X}$  (2)  ${}_6^{14}\text{X}$   
 (3)  ${}_6^{14}\text{X}$  (4)  ${}_6^{14}\text{X}$
6. The number of protons, neutrons and electrons in a neutral atom of the element  ${}_{10}^{24}\text{X}$  are \_\_\_\_\_ respectively.  
 (1) 89, 89, 242 (2) 89, 142, 89  
 (3) 89, 71, 89 (4) 89, 231, 89
7. Which of the following pairs represent isobars?  
 (1)  ${}_2^4\text{He}$  and  ${}_2^4\text{He}$  (2)  ${}_{12}^{24}\text{Mg}$  and  ${}_{12}^{24}\text{Mg}$   
 (3)  ${}_{19}^{40}\text{K}$  and  ${}_{20}^{40}\text{Ca}$  (4)  ${}_{19}^{40}\text{K}$  and  ${}_{19}^{39}\text{K}$
8. Rutherford's model of an atom is unstable because \_\_\_\_\_.  
 (1) the proton is an unstable particle  
 (2) the electron is an unstable particle  
 (3) the electron has very high kinetic energy  
 (4) an accelerated electron in an orbit emits radiation
9. The SI unit of frequency of radiation is  
 (1)  $\text{cm}^{-1}$  (2) Hertz  
 (3) nm (4) metre
10. \_\_\_\_\_ is the number of wavelengths per unit length.  
 (1) Amplitude (2) Velocity  
 (3) Wavenumber (4) Frequency
11. The value of Rydberg constant for hydrogen atom is \_\_\_\_\_.  
 (1)  $109677 \text{ m}^{-1}$  (2)  $109677 \text{ cm}^{-1}$   
 (3)  $109677 \times 10^7 \text{ cm}^{-1}$  (4)  $109677 \times 10^{-7} \text{ cm}^{-1}$
12. The emission spectrum of hydrogen in the ultraviolet region of electromagnetic radiation forms \_\_\_\_\_ series.  
 (1) Pfund (2) Lyman  
 (3) Balmer (4) Brackett
13. The spectral line obtained when an electron jumps from  $n = 6$  to  $n = 3$  level in hydrogen atom belongs to the \_\_\_\_\_ series.  
 (1) Balmer (2) Lyman  
 (3) Paschen (4) Pfund
14. Which of the following is INCORRECT regarding different series observed in hydrogen spectrum?  
 (1) For Balmer series,  $n_1 = 2$ .  
 (2) For Brackett series,  $n_1 = 4$ .  
 (3) For Lyman series,  $n_1 = 1$ .  
 (4) For Pfund series,  $n_1 = 3$
15. According to Bohr's model for the hydrogen atom, when electrons revolve in the stationary orbits, \_\_\_\_\_.  
 (1) the energies of the electrons remain constant  
 (2) they become stationary  
 (3) they gain potential energy  
 (4) there is increase in energies of the electrons

# Atom Structure Chemthink Answer Key

**Bud Morgan**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the Bud Morgan text.

## **Atom Structure Chemthink Answer Key:**

**ATOMIC STRUCTURE High School Chemistry Curriculum** Bud Morgan, 2025-07-09 Welcome to our comprehensive guide to Atomic Structure designed specifically for high school students studying chemistry In this book we will explore the fascinating world of atoms the basic building blocks of matter and delve into the complexities of their structure Chapter 1 provides a solid foundation by introducing the concept of atoms and their historical significance in the development of modern chemistry We will examine the key scientists who contributed to our understanding of atomic structure from Democritus to Thomson Rutherford and Bohr Moving on to Chapter 2 we will explore the structure of the atom in more detail including the subatomic particles that make up an atom protons neutrons and electrons We will also discuss the concept of atomic number atomic mass and isotopes In Chapter 3 we will delve into the arrangement of electrons within an atom s electron cloud including energy levels sublevels and orbital shapes We will also discuss the principles of electron configuration and how it impacts an atom s chemical properties Chapter 4 will focus on the periodic table and how it organizes elements based on their atomic structure We will explore periodic trends such as atomic radius ionization energy and electron affinity and discuss how these trends are related to an element s position on the table In our final chapter we will discuss the significance of atomic structure in various chemical reactions including bonding stoichiometry and the behavior of gases We will also explore the importance of understanding atomic structure in fields such as biochemistry environmental science and materials science Throughout this book we will provide clear explanations helpful illustrations and interactive activities to reinforce key concepts Whether you are a visual learner a hands on learner or a theoretical learner this book is designed to meet your needs We hope that this book will not only enhance your understanding of atomic structure but also inspire you to further explore the wonders of chemistry So let s dive in and unlock the mysteries of the atom together

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