

- **Pinched at bottom** of the cones.
- **Open**
 - **Delocalised π system** through the circles at bottom is the result
 - **Major input** for the **delocalised** due to **conjugated systems**
 - **then this π -electrons** is pushed through the **back-bondings** in the result too
 - the **delocalised π system** through the **circles**
 - **large delocalised double bonding system**

20. ALL properties of ionic, ionic/metal is polar

- **Water** is the **universal solvent** - has some for a **polar molecule**, meaning it has a **negative** and **positive charge** allowing it to **break apart** and **dissolve** substances
- **Order of matter**
 - **Solid**
 - **Liquid**
 - **Gas**
- **High level capacity** - matter is able to absorb high levels of heat **energy** without changing temperature

21. Elements of matter

- **Element** - are atoms forming a substance - or element
- **Atom** - is the smallest part of an element

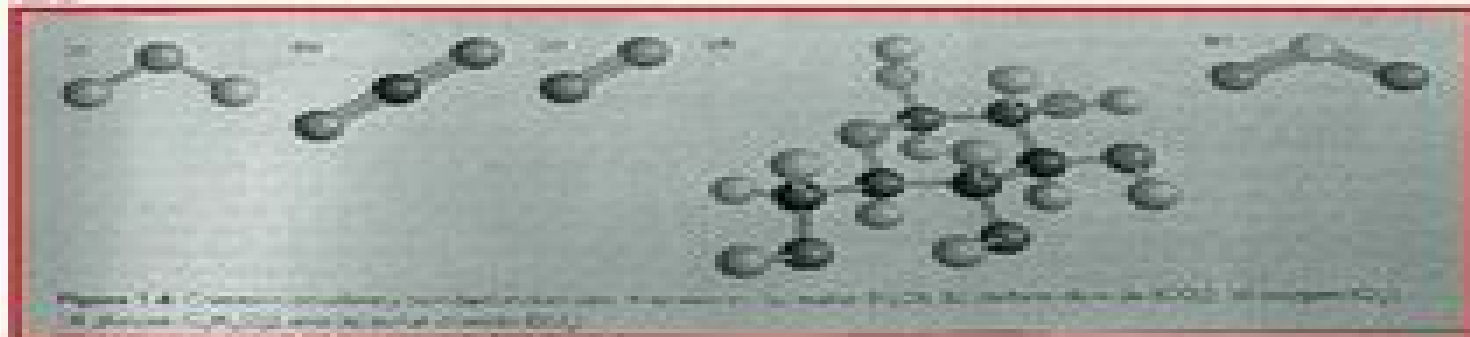
22. Positive, negative, electron (charge of each, where are electrons)

- **positive** - positive charge - found in the nucleus
- **negative** - negatively charged - found in the nucleus
- **electron** - negatively charged - orbit the nucleus

23. Compounds or molecules

- **Compound** - at least 2 different elements joined together through bonding (hydrogen and oxygen H_2O)
- **molecules** - where atoms are bonded - CO or CO_2

24. Point the molecule formulae for: 1st year, water, carbon dioxide, sulfur dioxide, $NaCl$, glucose, ethanol, amino acids, methane



25. Covalent bonds, ionic bonds, all hydrogen bonds

- **Ionic bonding** - A type of bond where atoms of opposite charge attract each other, relatively strong bond
 - **example** - Sodium Chloride ($NaCl$), $NaCl$ ions, Calcium Chloride ($CaCl_2$), Magnesium sulfate ($MgSO_4$)
- **Covalent bonding** - A type of bond where two atoms share electrons, both atoms have a complete shell (outer shell), so it is the strongest type of bond
 - **example** - Water (H_2O), Carbon Dioxide (CO_2), Oxygen (O_2), Ethane (C_2H_6), Ethene (C_2H_4), and Sulfur Dioxide (SO_2)
- **Hydrogen bonding** - relatively weak - occurs in the electron group molecules, where the size of the hydrogen bond is smaller, usually depend on how close the
 - **bonding of water molecules**

26. Define: Substrates, latent heat of vaporisation, latent heat of fusion, heat capacity, absorption

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