

Answers

AS Level

Topic 1: Chemical formulae and moles

Answers to in-text questions

Questions 1.1

- 1.5 mol
- 0.75 mol
- 0.667 mol

Questions 1.2

- 16.7 mol
- 8.3×10^{-3} mol

Questions 1.3

- 9
- 11
- 12
- 24
- 22

Questions 1.4

- 151.9
 - 162.1
 - 60.0
 - 132.1
 - 388.8
- 0.5 mol
 - 2.5 mol
 - 0.6 mol
 - 2.0 mol
 - 0.38 mol
- 180.6 g
 - 44.46 g

Questions 1.5

- $M_r = 277.9$ moles of oxygen = 11
- $M_r = 474.3$ moles of oxygen = 20

Questions 1.6

- Cu_2S
- C_3H_8
- CaFe_2O_4

Questions 1.7

- $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 - $\text{I}_2 + 3\text{Cl}_2 \rightarrow 2\text{ICl}_3$
 - $\text{NaOH} + \text{Al}(\text{OH})_3 \rightarrow \text{NaAlO}_2 + 2\text{H}_2\text{O}$
 - $2\text{H}_2\text{S} + \text{SO}_2 \rightarrow 3\text{S} + 2\text{H}_2\text{O}$
 - $4\text{NH}_3 + 3\text{O}_2 \rightarrow 2\text{N}_2 + 6\text{H}_2\text{O}$

- $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$
 - $\text{Pb(s)} + 2\text{AgNO}_3(\text{aq}) \rightarrow \text{Pb}(\text{NO}_3)_2(\text{aq}) + 2\text{Ag(s)}$
 - $\text{Na}_2\text{O(s)} + \text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)}$
 - $2\text{FeCl}_2 + \text{Cl}_2 \rightarrow 2\text{FeCl}_3$
 - $\text{Fe}_2(\text{SO}_4)_3 + 6\text{NaOH} \rightarrow 2\text{Fe}(\text{OH})_3 + 3\text{Na}_2\text{SO}_4$

Questions 1.8

- 17.0 g
- 60.7 g
- 51.1%

Questions 1.9

- 1.2 dm³
- Volume of $\text{SO}_2 = 50 \text{ cm}^3$, mass of sulfur = 0.20 g

Questions 1.10

- 2.67 mol dm⁻³
 - 0.025 mol dm⁻³
 - 0.833 mol dm⁻³
- 0.75 mol
 - 0.044 mol
 - 0.010 mol

Questions 1.11

- 1.2 dm³
- 12 dm³
- Volume of acid = 0.41 dm³, volume of $\text{H}_2 = 4.9 \text{ dm}^3$
- 0.64 g
- 1.27 mol dm⁻³

Answers to end of topic questions

- Any two of ZnCO_3 , $\text{Zn}(\text{OH})_2$ and ZnO
 - To make sure no water of crystallisation remained
 - Mass of zinc sulfate = mass of tube and salt after heating – mass of tube
 $= 76.34 - 74.25 = 2.09 \text{ g}$
 Relative molecular mass of zinc sulfate (ZnSO_4) = $65.4 + 32.1 + (4 \times 16.0) = 161.5$
 Amount in moles = $2.09 \div 161.5 = 0.01294 = 1.29 \times 10^{-2}$
 - Mass of water = mass of tube and hydrated salt – mass of tube and salt after heating
 $= 77.97 - 76.34 = 1.63 \text{ g}$
 Molar mass of water (H_2O) = $(2 \times 1) + 16 = 18$
 Amount in moles = $1.63 \div 18 = 0.0905 = 9.1 \times 10^{-2}$
 - Value of $x = 9.1 \times 10^{-2} \div 1.29 \times 10^{-2} = 7.054$
 which is 7 as a whole number

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L Darling-Hammond



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