

ATOMIC STRUCTURE

Exercises

Electrons in atoms are not particles, but quantum objects. It is not possible to find electrons by the usual measuring method. However, the observed pattern is useful for understanding the atomic structure. It depends on the electron's wave function. In calculating atoms, the electron's probability density is calculated. The electron's wave function is calculated with Schrödinger's equation. The electron's probability density is calculated with the wave function. The electron's probability density is calculated with the wave function.

The electron's wave function is calculated with the following equation:

$$\nabla^2 \psi + \frac{2m(E - V)}{\hbar^2} \psi = 0$$

where ψ is the wave function, ∇^2 is the Laplacian, m is the electron mass, E is the energy, and V is the potential energy.

Complete the following table:

Element	Atomic Number	Element Name	Group Number	Period	Electronegativity	Bohr Radius
H	1					
He	2					
Li	3					
Be	4					
B	5					
C	6					
N	7					
O	8					
F	9					
Ne	10					
Na	11					
Mg	12					
Al	13					
Si	14					
P	15					
S	16					
Cl	17					
Ar	18					

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