

CHAPTER 4 REVIEW

Arrangement of Electrons in Atoms**SECTION 2****SHORT ANSWER** Answer the following questions in the space provided.

1. _____ How many quantum numbers are used to describe the properties of electrons in atomic orbitals?
(a) 1 (c) 3
(b) 2 (d) 4
2. _____ A spherical electron cloud surrounding an atomic nucleus would best represent
(a) an s -orbital. (c) a combination of two different p -orbitals.
(b) a p -orbital. (d) a combination of an s and a p -orbital.
3. _____ How many electrons can an energy level of $n = 4$ hold?
(a) 32 (c) 8
(b) 24 (d) 6
4. _____ How many electrons can an energy level of $n = 2$ hold?
(a) 32 (c) 8
(b) 24 (d) 6
5. _____ Compared with an electron for which $n = 2$, an electron for which $n = 4$ has more
(a) spin. (c) energy.
(b) particle nature. (d) wave nature.
6. _____ According to Bohr, which is the point in the figure below where electrons cannot reside?
(a) point A (c) point C
(b) point B (d) point D



7. _____ According to the quantum theory, point D in the above figure represents:
(a) the fixed position of an electron.
(b) the furthest position from the nucleus that an electron can achieve.
(c) a position where an electron probably exists.
(d) a position where an electron cannot exist.

Arrangement Of Electrons In Atoms Chapter 4 Test Answers

Chris Tyreman



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Saraswati Chemistry Class 09 RP Manchanda, A text book on Chemistry *Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science*, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction **Test Bank to Accompany Chemistry and the Living Organism** Bloomfield, Catherine E.

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