

## 9.1

## THE ARITHMETIC OF EQUATIONS

## SECTION REVIEW

**Objectives**

- Calculate the amount of reactants required or product formed in a stoichiometric process.
- Interpret balanced chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

**Key Terms**

- stoichiometry

**Part A Completion**

Copy this completion exercise on blank paper and demonstrate your understanding of the concepts and terms it uses by completing the statements. Write down the complete word or phrase, short phrase, or name first.

- The coefficients in a balanced chemical equation indicate \_\_\_\_\_ 1. \_\_\_\_\_ the relative number of \_\_\_\_\_ 2. \_\_\_\_\_ of reactants and products. All stoichiometric calculations begin with a \_\_\_\_\_ 3. \_\_\_\_\_ (molar \_\_\_\_\_ 4. \_\_\_\_\_) and \_\_\_\_\_ 5. \_\_\_\_\_ are used to find the mass, moles, volume, and representative particles that react.
- In solving stoichiometric problems, conversion factors relating moles of reactants to \_\_\_\_\_ 6. \_\_\_\_\_ of products are used.
- If you assume \_\_\_\_\_ 7. \_\_\_\_\_, the equation also tells you about the volume of gas.

**Part B True-False**

(Circle) each of the statements to be true (write *True*) or false (write *False*), or correct (write *TC*) or incorrect (write *IT*).

- \_\_\_\_\_ 7. The coefficients in a balanced chemical equation can be used to find the relative molar masses of reacting reactants and products.
- \_\_\_\_\_ 8. The coefficients in a balanced chemical equation tell the relative numbers of reactants and products, expressed in any suitable units of volume.
- \_\_\_\_\_ 9. To calculate the mass of a molecule in grams, you can use the molar mass and Avogadro's number.

# Arithmetic Of Equations Section Answer Key

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