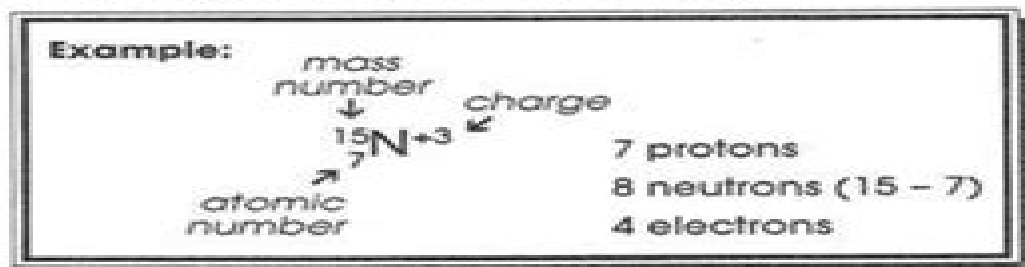


ATOMIC STRUCTURE

Name _____

An atom is made up of protons and neutrons (both found in the nucleus) and electrons (in the surrounding electron cloud). The atomic number is equal to the number of protons. The mass number is equal to the number of protons plus neutrons. In a neutral atom, the number of protons equals the number of electrons. The charge on an ion indicates an imbalance between protons and electrons. Too many electrons produces a negative charge, too few, a positive charge.

This structure can be written as part of a chemical symbol.



Complete the following chart.

| Element/ Ion | Atomic Number | Atomic Mass | Mass Number | Protons | Neutrons | Electrons |
|--|------------------|-------------|----------------|---------|----------|-----------|
| H | | | | | | |
| H ⁺ | | | | | | |
| ¹² ₆ C | | | | | | |
| ⁷ ₃ Li ⁺ | | | | | | |
| ³⁵ ₁₇ Cl ⁻ | | | | | | |
| ³⁹ ₁₉ K | | | | | | |
| ²⁴ ₁₂ Mg ²⁺ | | | | | | |
| As ³⁻ | | | | | | |
| Ag | | | | | | |
| Ag ⁺ | | | | | | |
| S ⁻² | | | | | | |
| U | | | | | | |

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Chemistry & Atomic Structure John David Main Smith, 1924 **Chemistry for Kids, the Atomic Structure** Lex Sharp, 2018-03-24 This book is meant for students grades 6 and higher It addresses concepts that are critical to the understanding of the inner structure of the atom which lays the foundation for appreciating how elements behave in chemical reactions The text has been integrated with colorful images that help engage the visual with the conceptual No prerequisites are necessary The book is suitable for the beginner who has not studied chemistry before and for the student who wants to clarify questions for review purposes A couple of mathematical chapters were also infused into the book to support younger learners with the necessary mathematical background to succeed To get the most from this book students should summarize each chapter and rehearse the concepts Attempting to work independently through all problems at the end of the book is beneficial only before reading the solutions When solving problems one can be sure the chemical principles presented were well understood Restricting oneself to only reading is incomplete learning The series walks learners gradually and conceptually through the chemical and particle physics theories that are meant to jumpstart a more profound scientific understanding Topics covered in this volume Subatomic Particle Atomic Mass Period Table of Elements Atomic Number Isotope and Ion Notations Atomic Weight The Mole Molar Mass Exponents The Rule of Three Practice problems and easy to follow solutions **Atomic Structure and Valency**, 1969 **ATOMIC STRUCTURE High School Chemistry**

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will delve into the arrangement of electrons within an atom's electron cloud including energy levels, sublevels, and orbital shapes. We will also discuss the principles of electron configuration and how it impacts an atom's chemical properties. Chapter 4 will focus on the periodic table and how it organizes elements based on their atomic structure. We will explore periodic trends such as atomic radius, ionization energy, and electron affinity, and discuss how these trends are related to an element's position on the table. In our final chapter, we will discuss the significance of atomic structure in various chemical reactions, including bonding, stoichiometry, and the behavior of gases. We will also explore the importance of understanding atomic structure in fields such as biochemistry, environmental science, and materials science. Throughout this book, we will provide clear explanations, helpful illustrations, and interactive activities to reinforce key concepts. Whether you are a visual learner, a hands-on learner, or a theoretical learner, this book is designed to meet your needs. We hope that this book will not only enhance your understanding of atomic structure but also inspire you to further explore the wonders of chemistry. So let's dive in and unlock the mysteries of the atom together.

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