

# Answers

## AS Level

### Topic 1: Chemical formulae and moles

#### Answers to in-text questions

##### Questions 1.1

- 1.5 mol
- 0.75 mol
- 0.667 mol

##### Questions 1.2

- 16.7 mol
- $8.3 \times 10^{-3}$  mol

##### Questions 1.3

- 9
- 11
- 12
- 24
- 22

##### Questions 1.4

- 151.9
  - 162.1
  - 60.0
  - 132.1
  - 388.8
- 0.5 mol
  - 2.5 mol
  - 0.6 mol
  - 2.0 mol
  - 0.38 mol
- 180.6 g
  - 44.46 g

##### Questions 1.5

- $M_r = 277.9$  moles of oxygen = 11
- $M_r = 474.3$  moles of oxygen = 20

##### Questions 1.6

- $\text{Cu}_2\text{S}$
- $\text{C}_3\text{H}_8$
- $\text{CaFe}_2\text{O}_4$

##### Questions 1.7

- $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
  - $\text{I}_2 + 3\text{Cl}_2 \rightarrow 2\text{ICl}_3$
  - $\text{NaOH} + \text{Al}(\text{OH})_3 \rightarrow \text{NaAlO}_2 + 2\text{H}_2\text{O}$
  - $2\text{H}_2\text{S} + \text{SO}_2 \rightarrow 3\text{S} + 2\text{H}_2\text{O}$
  - $4\text{NH}_3 + 3\text{O}_2 \rightarrow 2\text{N}_2 + 6\text{H}_2\text{O}$

- $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$
  - $\text{Pb(s)} + 2\text{AgNO}_3(\text{aq}) \rightarrow \text{Pb}(\text{NO}_3)_2(\text{aq}) + 2\text{Ag(s)}$
  - $\text{Na}_2\text{O(s)} + \text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)}$
  - $2\text{FeCl}_3 + \text{Cl}_2 \rightarrow 2\text{FeCl}_4$
  - $\text{Fe}_2(\text{SO}_4)_3 + 6\text{NaOH} \rightarrow 2\text{Fe}(\text{OH})_3 + 3\text{Na}_2\text{SO}_4$

##### Questions 1.8

- 17.0 g
- 60.7 g
- 51.1%

##### Questions 1.9

- 1.2 dm<sup>3</sup>
- Volume of  $\text{SO}_2 = 50 \text{ cm}^3$ , mass of sulfur = 0.20 g

##### Questions 1.10

- 2.67 mol dm<sup>-3</sup>
  - 0.025 mol dm<sup>-3</sup>
  - 0.833 mol dm<sup>-3</sup>
- 0.75 mol
  - 0.044 mol
  - 0.010 mol

##### Questions 1.11

- 1.2 dm<sup>3</sup>
- 12 dm<sup>3</sup>
- Volume of acid = 0.41 dm<sup>3</sup>, volume of  $\text{H}_2 = 4.9 \text{ dm}^3$
- 0.64 g
- 1.27 mol dm<sup>-3</sup>

#### Answers to end of topic questions

- Any two of  $\text{ZnCO}_3$ ,  $\text{Zn}(\text{OH})_2$  and  $\text{ZnO}$
  - To make sure no water of crystallisation remained
    - Mass of zinc sulfate = mass of tube and salt after heating – mass of tube  
 $= 76.34 - 74.25 = 2.09 \text{ g}$   
 Relative molecular mass of zinc sulfate ( $\text{ZnSO}_4$ ) =  $65.4 + 32.1 + (4 \times 16.0) = 161.5$   
 Amount in moles =  $2.09 \div 161.5 = 0.01294 = 1.29 \times 10^{-2}$
    - Mass of water = mass of tube and hydrated salt – mass of tube and salt after heating  
 $= 77.97 - 76.34 = 1.63 \text{ g}$   
 Molar mass of water ( $\text{H}_2\text{O}$ ) =  $(2 \times 1) + 16 = 18$   
 Amount in moles =  $1.63 \div 18 = 0.0905 = 9.1 \times 10^{-2}$
    - Value of  $x = 9.1 \times 10^{-2} \div 1.29 \times 10^{-2} = 7.054$   
 which is 7 as a whole number

# A Level Questions And Answers Chemistry A Level Questions And Answers Series

**Robert R Harr**



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