

Name: \_\_\_\_\_ Period: \_\_\_\_\_ Date: \_\_\_\_\_

## Student Exploration: Boyle's Law and Charles' Law

**Essential** questions are thought-provoking questions that guide your learning.

**Essential Question:** How does Boyle's Law and Charles' Law work?

As you explore these questions, you will learn how Boyle's Law and Charles' Law work. You will also learn how Boyle's Law and Charles' Law are related to the Ideal Gas Law.

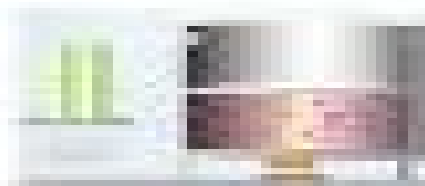
The goal of this exploration is to help you understand Boyle's Law and Charles' Law. You will learn how Boyle's Law and Charles' Law work. You will also learn how Boyle's Law and Charles' Law are related to the Ideal Gas Law.

### Boyle's Law

Boyle's Law states that the pressure of a gas is inversely proportional to its volume. This means that if the volume of a gas increases, the pressure will decrease. Conversely, if the volume of a gas decreases, the pressure will increase.

1. Which of the following graphs best represents Boyle's Law?

Graph A \_\_\_\_\_ Graph B \_\_\_\_\_



2. Which of the following graphs best represents Charles' Law?

Graph A \_\_\_\_\_ Graph B \_\_\_\_\_

Graph A shows a linear relationship where pressure increases as volume increases. Graph B shows an inverse relationship where pressure decreases as volume increases. Charles' Law states that the volume of a gas is directly proportional to its temperature. This means that if the temperature of a gas increases, the volume will also increase. Conversely, if the temperature of a gas decreases, the volume will also decrease.

3. Which of the following graphs best represents the Ideal Gas Law?

Graph A \_\_\_\_\_ Graph B \_\_\_\_\_

Graph A shows a linear relationship where pressure increases as volume increases. Graph B shows an inverse relationship where pressure decreases as volume increases. The Ideal Gas Law states that the pressure of a gas is directly proportional to its temperature and inversely proportional to its volume. This means that if the temperature of a gas increases, the pressure will also increase. Conversely, if the temperature of a gas decreases, the pressure will also decrease. Additionally, if the volume of a gas increases, the pressure will decrease. Conversely, if the volume of a gas decreases, the pressure will increase.

# Boyles Law Gizmo Answers Explore Learning

**RS Peters**



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