

# Bond Energy

What makes a reaction endothermic or exothermic?

## Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** ( $\Delta H$ ). In this activity we will look at one way energy changes can be approximated for chemical reactions.

## Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )		Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
  - When bonds are (broken/formed) there is a positive energy change.
  - Breaking bonds is (endothermic/exothermic).
  - When bonds are (broken/formed) there is a negative energy change.
  - Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.

- How do the changes in energy for the reverse reactions compare?

*The changes in energy for these reactions are the same, but opposite in sign.*

- Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.

*If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.*

# Bond Energy Pogil Key Extension Questions

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1837 and the Congressional Globe 1833 1873      **Chemistry in Context** Conrad L. Stanitski,1999      *Dynamics of Molecules and Chemical Reactions* Robert Wyatt,1996-06-27 Covers both molecular and reaction dynamics The work presents important theoretical and computational approaches to the study of energy transfer within and between molecules discussing the application of these approaches to problems of experimental interest It also describes time dependent and time independent methods variational and perturbative techniques iterative and direct approaches and methods based upon the use of physical grids of finite sets of basis function      **Catalog of Copyright Entries. Third Series** Library of Congress. Copyright Office,1960 Includes Part 1 Number 2 Books and Pamphlets Including Serials and Contributions to Periodicals July December      *From Quantum Mechanics to Biochemistry* ,1991      Papers ,1965      *The End of Religion, the Beginning of Self* Bruno M. Curfs,Bruno Curfs M.Sc.,2012-04 For your convenience there is a dedicated website available [click here](#) Also connect on Facebook With The End Of Religion The Beginning Of Self readers can look forward to an in depth analysis profound interpretation and insightful reflection of the Bible This spectacular read freely shares a thought provoking perspective about famous and less famous Bible stories and their practical significance in the reader's life Every letter of the Hebrew Alphabet is explained in a separate chapter and in many different ways as a letter and as a letter name as a number and as a number name as part of scripture and as part of words This technique is one of the most important legacies given to human kind There is no doubt scientists will be baffled by this innovative knowledge for centuries to come especially when they get wind of its applications Exhilarating this read contains poems riddles and appendices for extra depth and emphasis So what are you waiting for Let this book open the doors to a greater and more profound understanding about life the universe and everything      Chemistry insights 'O' level Rex M. Heyworth,2007

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