

# Solutions Manual for Fluid Mechanics 5th Edition by Kundu Cohen and Dowling

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## Chapter 1 Introduction

**1.1** A gas at 20°C may be *rarefied* if it contains less than  $10^{12}$  molecules per  $\text{mm}^3$ . If Avogadro's number is  $6.023\text{E}23$  molecules per mole, what air pressure does this represent?

**Solution:** The mass of one molecule of air may be computed as

$$m = \frac{\text{Molecular weight}}{\text{Avogadro's number}} = \frac{28.97 \text{ mol}^{-1}}{6.023\text{E}23 \text{ molecules/g} \cdot \text{mol}} = 4.81\text{E}-23 \text{ g}$$

Then the density of air containing  $10^{12}$  molecules per  $\text{mm}^3$  is, in SI units,

$$\begin{aligned} \rho &= \frac{(10^{12} \text{ molecules})}{\text{mm}^3} \cdot \frac{4.81\text{E}-23 \text{ g}}{\text{molecule}} \\ &= 4.81\text{E}-11 \frac{\text{g}}{\text{mm}^3} = 4.81\text{E}-5 \frac{\text{kg}}{\text{m}^3} \end{aligned}$$

Finally, from the perfect gas law, Eq. (1.13), at 20°C = 293 K, we obtain the pressure:

$$p = \rho RT = 4.81\text{E}-5 \frac{\text{kg}}{\text{m}^3} \cdot \frac{\text{m}^2}{\text{s}^2 \cdot \text{K}} (293 \text{ K}) = 4.0 \text{ Pa} \quad \text{Ans.}$$

**1.2** The earth's atmosphere can be modeled as a uniform layer of air of thickness 20 km and average density  $0.6 \text{ kg/m}^3$  (see Table A-6). Use these values to estimate the total mass and total number of molecules of air in the entire atmosphere of the earth.

**Solution:** Let  $R_e$  be the earth's radius  $\approx 6377 \text{ km}$ . Then the total mass of air in the atmosphere is

$$\begin{aligned} m_a &= \rho \, d\text{Vol} = \rho_{\text{avg}} (\text{Air Vol}) \approx \rho_{\text{avg}} 4\pi R_e^2 (\text{Air thickness}) \\ &= (0.6 \text{ kg/m}^3) 4\pi (6.377\text{E}6 \text{ m})^2 (20\text{E}3 \text{ m}) \approx 6.1\text{E}18 \text{ kg} \quad \text{Ans.} \end{aligned}$$

Dividing by the mass of one molecule  $\approx 4.8\text{E}-23 \text{ g}$  (see Prob. 1.1 above), we obtain the total number of molecules in the earth's atmosphere:

$$N_{\text{molecules}} = \frac{m(\text{atmosphere})}{m(\text{one molecule})} = \frac{6.1\text{E}21 \text{ grams}}{4.8\text{E}-23 \text{ gm/molecule}} \approx 1.3\text{E}44 \text{ molecules} \quad \text{Ans.}$$

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**Terry C. Jones**



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