Chemistry 338 2nd Semester Final Exam Review Sheet

Unit 1 - Matter & Measurement - Chapters 1 & 2

 A student experimentally determines the density of an object to be 4.2 g/mL. If the accepted value is 5.0 g/mL, calculate the % error.

% error = $\frac{experimentat-accepted}{accepted} \times 100$ % error = $\frac{4.2-5.0}{5.0} \times 100 = -16\%$ error

Unit 2 - Periodic Table - Chapter 5

2.

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass#	Charge
Fungsten ion	194 W+2	7-4	74	110	72	184	+2

 Using the following atomic radius values, predict the atomic radius of potassium (K): Li = 145 pm, Na = 180 pm, K = 7, Rb = 235 pm K = 220 pm

Unit 3 - Reactions/Compounds - Chapters 7 & 8

- 4. Classify the compound as ionic or covalent. Name the compound
 - Fe₂O₃ iron (III) oxide (ionic roman numerals for transition metals, change ending to -ide)
 - b. Cl₂O₈ dichlorine octaoxide (covalent use prefixes, change ending to -ide)
- 5. Classify the compound as ionic or covalent. Write the formula
 - a. Trisulfur Dinitride S₃N₂ (covalent)
 - b. Zinc Hydroxide Zn(OH)2 (ionic look up ion charges and criss-cross) Zn⁻² OH⁻³
 - c. Chromium (III) nitrate Cr(NO₃)₃ (ionic criss-cross) Cr²³ NO₃⁻¹
- 6. Write the equation, balance it, and classify the type of reaction:

Iron (II) sulfate reacts with sodium hydroxide to yield iron(II) hydroxide and sodium sulfate

> 1 FeSO₄ + 2 NaOH → 1 Fe(OH)₂ + 1 Na₂SO₄ double displacement (two ionic compounds reacting)

7. Balance the equation and classify the type of reaction:

1 Cu + 2 AgNO₃ → 1 Cu(NO₃)₂ + 2 Ag Single displacement (element and a compound reacting)

Unit 4 - Moles - Chapter 3

8. Calculate the molar mass of Fe₂(SO₄)₅.

Fe (2)(55.9 g/mol) = 111.8 g/mol

S (3)(32.1 g/mol) = 96.3 g/mol

+ O (12)(16.0 g/mol) = +192 g/mol

= 400.1 g/mol

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Basic Inorganic Chemistry F. Albert Cotton, Geoffrey Wilkinson, Paul L. Gaus, 1995-01-18 Explains the basics of inorganic chemistry with a primary emphasis on facts then uses the student's growing factual knowledge as a foundation for discussing the important principles of periodicity in structure bonding and reactivity New to this updated edition improved treatment of atomic orbitals and properties such as electronegativity novel approaches to the depiction of ionic structures nomenclature for transition metal compounds quantitative approaches to acid base chemistry Wade's rules for boranes and carboranes the chemistry of major new classes of substances including fullerenes and silenes plus a chapter on the inorganic Catalog of Copyright Entries. Third Series Library of Congress. Resources in Education ,1997-10 solid state Copyright Office, 1967 Includes Part 1 Number 1 Books and Pamphlets Including Serials and Contributions to Periodicals REA's Authoritative Guide to Medical & Dental Schools Research and Education January June Association, 1996-12-01 An excellent source book for those who are beginning the medical or dental school application process Included are profiles on every U S And Canadian medical and dental school as well as information on select foreign medical schools Also included are sections on osteopathic schools chiropractic schools and podiatric schools Important information is also included on undergraduate preparation the application process financial aid and graduation requirements

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Nursing School Entrance Exam Review William H. Brown, Frederick A. Bettelheim, Jerry March, 2004-04 Curriculum Guides for the Science Department of the Highland Park Public Schools for Grades 9-12 Highland Park School District, 1958 Vertical File Service Catalog, 1947

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