Chemistry 338 2nd Semester Final Exam Review Sheet

Unit 1 - Matter & Measurement - Chapters 1 & 2

 A student experimentally determines the density of an object to be 4.2 g/mL. If the accepted value is 5.0 g/mL, calculate the % error.

% error = $\frac{experimentat-accepted}{accepted} \times 100$ % error = $\frac{4.2-5.0}{5.0} \times 100 = -16\%$ error

Unit 2 - Periodic Table - Chapter 5

2.

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass#	Charge
Fungsten ion	194 W+2	7-4	74	110	72	184	+2

 Using the following atomic radius values, predict the atomic radius of potassium (K): Li = 145 pm, Na = 180 pm, K = 7, Rb = 235 pm K = 220 pm

Unit 3 - Reactions/Compounds - Chapters 7 & 8

- 4. Classify the compound as ionic or covalent. Name the compound
 - Fe₂O₃ iron (III) oxide (ionic roman numerals for transition metals, change ending to -ide)
 - b. Cl₂O₈ dichlorine octaoxide (covalent use prefixes, change ending to -ide)
- 5. Classify the compound as ionic or covalent. Write the formula
 - a. Trisulfur Dinitride S₃N₂ (covalent)
 - b. Zinc Hydroxide Zn(OH)2 (ionic look up ion charges and criss-cross) Zn⁻² OH⁻³
 - c. Chromium (III) nitrate Cr(NO₃)₃ (ionic criss-cross) Cr²³ NO₃⁻¹
- 6. Write the equation, balance it, and classify the type of reaction:

Iron (II) sulfate reacts with sodium hydroxide to yield iron(II) hydroxide and sodium sulfate

> 1 FeSO₄ + 2 NaOH → 1 Fe(OH)₂ + 1 Na₂SO₄ double displacement (two ionic compounds reacting)

7. Balance the equation and classify the type of reaction:

1 Cu + 2 AgNO₃ → 1 Cu(NO₃)₂ + 2 Ag Single displacement (element and a compound reacting)

Unit 4 - Moles - Chapter 3

8. Calculate the molar mass of Fe₂(SO₄)₅.

Fe (2)(55.9 g/mol) = 111.8 g/mol

S (3)(32.1 g/mol) = 96.3 g/mol

+ O (12)(16.0 g/mol) = +192 g/mol

= 400.1 g/mol

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