

## 2 - Bonding & Molecular Structures

10 20 30 40 50 60 70 80 90 100

Table 4.1 - Some Average Single and Multiple Bond Strengths (kJ/mol)

	H	C	N	O	F	Cl	Br	I	S	Se	Te
H	436	413	391	460	565	242	233	213	364	267	236
C		347	305	336	485	337	276	242	272	226	200
N			391	347	460	312	259	228	272	226	200
O				498	500	337	276	242	364	267	236
F					485	312	259	228	272	226	200
Cl						337	276	242	364	267	236
Br							276	242	364	267	236
I								242	364	267	236
S									364	267	236
Se										267	236
Te											236

Multiple Bonds

C=C	614	C≡C	839
C=O	799	C≡O	1076
C=N	891	C≡N	960
C=O and N=O	608	C≡O and N≡N	945

Table 4.2 - Standard Enthalpies of Formation (kJ/mol)

Compound	Standard Enthalpy of Formation	Compound	Standard Enthalpy of Formation
H <sub>2</sub> (g)	0	H <sub>2</sub> O(l)	-285.8
O <sub>2</sub> (g)	0	CO <sub>2</sub> (g)	-393.5

1. Write the balanced chemical equation for the complete combustion of ethane (C<sub>2</sub>H<sub>6</sub>).



2. Draw structural formulae for each reactant and product species.



3. Calculate the standard enthalpy change for the reaction. 111.5 kJ/mol  
 Calculate the standard enthalpy change for the reaction. 111.5 kJ/mol

4. What is the standard enthalpy of formation for CO<sub>2</sub>(g)? -393.5 kJ/mol

5. Use the bond enthalpies given in the table to calculate the standard enthalpy of formation for CO<sub>2</sub>(g).

# Bonding And Molecular Structure Worksheet Answers South Pasadena

**George E. Ryschkewitsch**



## Bonding And Molecular Structure Worksheet Answers South Pasadena:

Chemical Bonds Harry B. Gray, 1994-12-05 This profusely illustrated book by a world renowned chemist and award winning chemistry teacher provides science students with an introduction to atomic and molecular structure and bonding This is a reprint of a book first published by Benjamin Cummings 1973      Structure and Bonding Jack Barrett, 2001 This book explains in non mathematical terms where possible the factors that govern covalent bond formation the lengths and strengths of bonds and molecular shapes      Chemical Structure and Bonding Roger L. DeKock, Harry B. Gray, 1989 Designed for use in inorganic physical and quantum chemistry courses this textbook includes numerous questions and problems at the end of each chapter and an Appendix with answers to most of the problems      *Chemical Bonding and the Geometry of Molecules* George E. Ryschkewitsch, 1965      **Chemical Bonding and the Geometry of Molecules** George Eugene Ryschkewitsch, 1965      *The Shape and Structure of Molecules* Charles Alfred Coulson, 1982      **Bonding and Molecular Structure** Russell S. Drago, 1971      *A Pictorial Approach to Molecular Bonding* John G. Verkade, 2012-12-06 With the development of accurate molecular calculations in recent years useful predictions of molecular electronic properties are currently being made It is therefore becoming increasingly important for the non theoretically oriented chemist to appreciate the underlying principles governing molecular orbital formation and to distinguish them from the quantitative details as associated with particular molecules It seems highly desirable then that the non theoretician be able to deduce results of general validity without esoteric mathematics In this context pictorial reasoning is particularly useful Such an approach is virtually indispensable if bonding concepts are to be taught to chemistry students early in their careers Undergraduate chemistry majors typically find it difficult to formulate molecular orbital schemes especially delocalized ones for molecules more complicated than diatomics The major reason for this regrettable situation is the general impracticability of teaching group theory before students take organic and inorganic courses wherein the applications of these concepts are most beneficial Consequently many students graduate with the misconception that the ground rules governing bonding in molecules such as  $\text{NH}_3$  are somehow different from those which apply to aromatic systems such as  $\text{C}_6\text{H}_6$  Conversely seniors and many graduate students are usually only vaguely if at all aware that sigma bonding like extended pi bonding can profitably be described in a delocalized manner when discussing the UV photoelectron spectrum of  $\text{CH}_4$  for example      *The Nature of the Chemical Bond and the Structure of Molecules and Crystals* Linus Pauling, 1963      The Theory of Molecular Structure and Bonding, 1980      Molecular Structure and Conformation I. G. Csizmadia, 1982 Good No Highlights No Markup all pages are intact Slight Shelfwear may have the corners slightly dented may have slight color changes slightly damaged spine      **Nature of the Chemical Bond and the Structure of Molecules and Crystals** Linus Pauling, 1967      *The Determination of Molecular Structure* P.J. Wheatley, 1981      **Structure of Molecules and the Chemical Bond** I. A. Kivovich Syrkin, M. E. Diatkina, M. E. Diatkina, 1950      The Determination of Molecular Structure Peter Jaffrey

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