

# Bond Energy

What makes a reaction endothermic or exothermic?

## Why?

Chemical reactions can be a lot like playing with Legos<sup>®</sup>—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the heat of reaction or the change in enthalpy ( $\Delta H$ ). In this activity we will look at one way energy changes can be approximated for chemical reactions.

## Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol)		Reaction	Change in Energy (kJ/mol)
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
  - When bonds are (broken)formed) there is a positive energy change.
  - Breaking bonds is (endothermic)exothermic).
  - When bonds are (broken)formed) there is a negative energy change.
  - Forming bonds is (endothermic)exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.
  - How do the changes in energy for the reverse reactions compare?  
*The changes in energy for these reactions are the same, but opposite in sign.*
  - Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.  
*If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.*

# Bond Energy Graph Pogil

**Daniela Niemeyer**

## **Bond Energy Graph Pogil:**

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**Introduction to Bond Graphs and their Applications** Jean U. Thoma, 2016-05-04 Introduction to Bond Graphs and Their Applications is an introductory text on bond graphs and their applications in the field of engineering The applications of bond graphs in mechanical engineering and design fluid mechanics electronic data processing and thermal and thermodynamic systems are discussed This book is comprised of eight chapters and begins by comparing the different kinds of graphs diagrams and models before turning to the fundamentals of bond graphs The next chapter introduces the reader to the systematic application of bond graphs in mechanical engineering and design fluid power engineering sometimes called oil hydraulics electrotechnique and electronics and thermodynamics The use of bond graphs in automatic computer programming with the ENPORT program is also described The final chapter is devoted to inertia and resistance fields linear two ports in different causalities thermodynamics of flow processes electromechanical components systems with distributed parameters and force and velocity as effort or flow This monograph is intended primarily for all engineers interested in representing simple or complex engineering systems and should also be of value to students in the different engineering disciplines mechanics fluid mechanics and electronics with electromechanical power conversion or thermodynamics

**Bond Graph Modelling of Exergy in Integrated Energy Systems** Robin Wardle, 2022 *Chemical Bonds and Bonds Energy* Robert Thomas Sanderson, 1976-06-28 Chemical Bonds and Bonds Energy Second Edition provides information pertinent to the fundamental aspects of contributing bond energy and bond dissociation energy This book explores the values that are useful in the interpretation of significant phenomena such as product distribution and reaction

mechanisms Organized into 12 chapters this edition begins with an overview of the quantitative relationship among three basic properties of an atom namely nonpolar covalent radius electronegativity and homonuclear single covalent bond energy This text then examines the quantitative means of evaluating the partial atomic charges that result from initial differences in the electromagnetivity of atoms that form a compound Other chapters consider the recognition of the reduction of bond weakening not by multiplicity and in certain types of single covalent bonds The final chapter deals with the application of the principal ideas and techniques to the oxidation of ethane This book is a valuable resource for organic and inorganic chemists

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Universitetscenter. Institut for Studiet af Matematik og Fysik samt deres Funktioner i Undervisning, Forskning og Anvendelser,2005    **Energy Bond Graphs** ,2003    Bond Graphs for Causal Explanations J. L. Top,1992    **The Chemical Bond in Inorganic Chemistry** Ian David Brown,2002 there has long been a need for a dedicated monograph on the subject a highly readable book about a theory that though it has long found application in inorganic crystal chemistry deserves to be used more widely Crystallography NewsThe bond valence model is a recently developed model of the chemical bond in inorganic chemistry that complements the bond model widely used in organic chemistry It is simple quantitative intuitive and predictive no more than a pocket calculator is needed to calculate it This book focuses on the theory that underlies the model and shows how it has been used in physics materials science chemistry mineralogy soil science and molecular biology    **Bond Graph Techniques for Dynamic Systems in Engineering and Biology** Dean Karnopp,1979

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