

5 - Bonding & Molecular Structures

10 20 30 40 50 60 70 80 90 100

Table 4.4 - Some Average Single and Multiple Bond Strengths (kJ/mol)

	H	C	N	O	F	Cl	Br	I	S	Se	Te
H	436										
C		413	347	336	485						
N			391	305	460						
O				443	459						
F					485						
Cl						328					
Br							276				
I								240			
S									265		
Se										226	
Te											226

Multiple Bonds

C=C	614	C≡C	839
C=O	799	C≡O	1072
C=N	891	C≡N	960
C=O and N=O	607	C≡O and N≡N	945

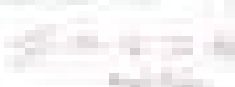
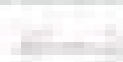
Table 4.5 - Standard Enthalpies of Formation (kJ/mol)

Compound	Standard Enthalpy of Formation	Compound	Standard Enthalpy of Formation
$\text{H}_2\text{O}(l)$	-285.8	$\text{H}_2\text{O}(g)$	-241.8
$\text{CO}_2(g)$	-393.5	$\text{CO}(g)$	-137.2

1. Write the balanced chemical equation for the complete combustion of ethane. (10pts)



2. Draw structural formula for each molecule and label the geometry. (10pts)



3. Calculate the enthalpy change for the reaction of the reactants.
 Calculate the enthalpy change for the reaction of the products. (10pts)



4. What is the $\Delta H_{\text{combustion}}$ for the reaction? (10pts)

5. Use the bond enthalpies provided to calculate the $\Delta H_{\text{combustion}}$ for the reaction. (10pts)

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