

# Bond Energy

What makes a reaction endothermic or exothermic?

## Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** ( $\Delta H$ ). In this activity we will look at one way energy changes can be approximated for chemical reactions.

## Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )		Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
  - When bonds are (broken/formed) there is a positive energy change.
  - Breaking bonds is (endothermic/exothermic).
  - When bonds are (broken/formed) there is a negative energy change.
  - Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.
  - How do the changes in energy for the reverse reactions compare?
 

*The changes in energy for these reactions are the same, but opposite in sign.*
  - Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.
 

*If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.*

# Bond Energy Pogil Extension Questions

**Tom Ziegler**



**Bond Energy Pogil Extension Questions:**

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This book delves into Bond Energy Pogil Extension Questions. Bond Energy Pogil Extension Questions is a vital topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Bond Energy Pogil Extension Questions, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
  - Chapter 1: Introduction to Bond Energy Pogil Extension Questions
  - Chapter 2: Essential Elements of Bond Energy Pogil Extension Questions
  - Chapter 3: Bond Energy Pogil Extension Questions in Everyday Life
  - Chapter 4: Bond Energy Pogil Extension Questions in Specific Contexts
  - Chapter 5: Conclusion
2. In chapter 1, the author will provide an overview of Bond Energy Pogil Extension Questions. This chapter will explore what Bond Energy Pogil Extension Questions is, why Bond Energy Pogil Extension Questions is vital, and how to effectively learn about Bond Energy Pogil Extension Questions.
3. In chapter 2, this book will delve into the foundational concepts of Bond Energy Pogil Extension Questions. This chapter will elucidate the essential principles that need to be understood to grasp Bond Energy Pogil Extension Questions in its entirety.
4. In chapter 3, the author will examine the practical applications of Bond Energy Pogil Extension Questions in daily life. This chapter will showcase real-world examples of how Bond Energy Pogil Extension Questions can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Bond Energy Pogil Extension Questions in specific contexts. This chapter will explore how Bond Energy Pogil Extension Questions is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Bond Energy Pogil Extension Questions. This chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Bond Energy Pogil Extension Questions.

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