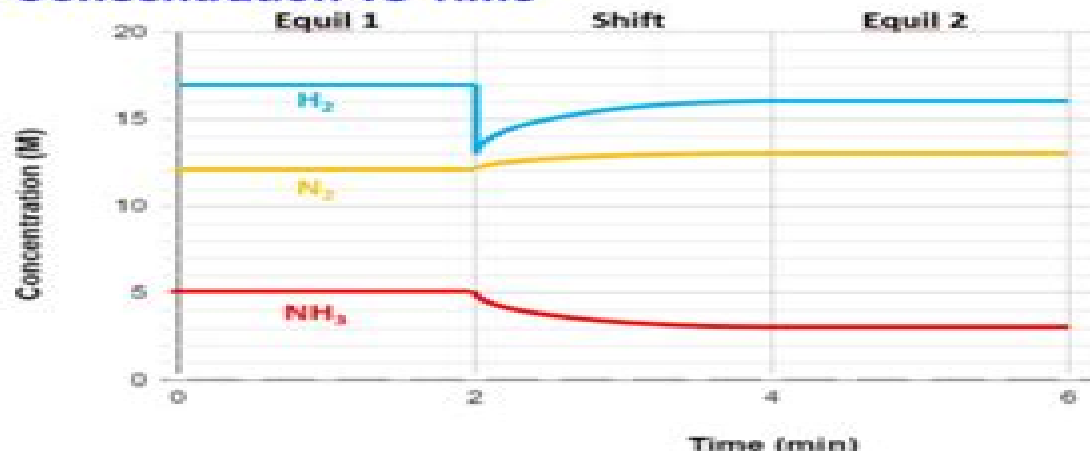


## Inquiry Question

## An Exercise in Stress - Applying Le Chatelier's Principle (Lab Analysis)

Name: \_\_\_\_\_ Date: \_\_\_\_\_

## Concentration vs Time



Example equilibrium:



From the graph above, it can be seen that the system is at equilibrium for the first two minutes. Then at the 2min mark a stress is applied (some H<sub>2</sub> is removed). Following that the equilibrium shifts left, which is indicated by the curved sections. Note that the height of the curves is in a 1:3:2 ratio, which matches the mole ratio in the equation. After 4 minutes the system has re-established equilibrium. Ideally the K<sub>eq</sub> value calculated from E1 and E2 would be the same, but there are often sources of error, and other factors, in lab situations.

Le Chatelier's Principle tells us that when an equilibrium system is stressed, the reaction will shift to reduce the effect of the stress and try to re-establish equilibrium. Changes in pressure, volume, concentration, and temperature offer "stresses" that create a disturbance to the system. The goal of this lab is to apply your knowledge of Le Chatelier's Principle and equilibrium to interpret a Concentration vs Time graph. You will use the graph and predict the stress that occurred to cause the changes in the graph. You will also calculate K<sub>eq</sub> based on equilibrium concentrations read off the graph.

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