

AQA Science - Covalent bonding

1. Identify the **group numbers** of the elements involved
2. The number of electrons needed to become stable is the number of **'pairs'** in the shared area (each pair should be made of one dot and one cross)
3. Fill in the **shared area** with the pairs
4. The **electrons remaining** from the original atom are placed on the **outer shell**.

Worked example 1: Chlorine - Cl_2

1. Chlorine is in group 7
2. Chlorine needs 1 electron to become stable so there will be 1 pair of electrons in the shared area
3. Fill in the shared area with the 1 pair
4. There are 6 electrons left so place these on the outer shell



Worked example 2: water H_2O

1. Hydrogen has 1 electron in the 1st shell. Oxygen is in group 6
2. Hydrogen needs 1 electron to become stable and oxygen needs 2 so there will be 2 hydrogen atoms
3. Fill in the shared areas with 1 pair in each
4. The only hydrogen electron is used so there are no others to place. There are 4 electrons left on the oxygen atom and so these are placed on the outer shell



Your turn

Hydrochloric acid - HCl

Ammonia - NH_3

Methane - CH_4

Carbon Dioxide - CO_2

Oxygen - O_2

Nitrogen - N_2

Core Teaching Resources Chemistry Test 8 Covalent Bonding

Mike Jess



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