Types of Chemical Reactions Worksheet #1

HJ/Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 2Mg + O₂ → 2 MgO Synthesis
- Cal₂ + Cl₂ → CaCl₂ + I₂ Single Replacement
- 3) 3KOH + AlCl₃ → Al(OH)₃ + 3KCl Double Replacement
- C + O₂ → CO₂ Synthesis
- *5) Ca(OH)₂ + 2HCl → CaCl₂ + 2H₂O Double Replacement
- 6) 2KClO₃ → 2KCl+3 O₂ Decomposition
- 3CuSO₄ + 2Al → Al₂(SO₄)₃ + 3Cu Single Replacement
- 8) Na₂S + 2AgNO₃ → 2NaNO₃ + Ag₂S Double Replacement
- 9) 3H₂ + N₂ → 2NH₃ Synthesis
- *10) 2Na + 2H2O -> 2NaOH + H2 Single Replacement

*Remember: "H2O" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- Double replacement: Na₂CrO₄ + PbCl₂ → 2 NaCl + PbCrO₄
- Single replacement: Cl₂ + 2 NaBr → 2 NaCl + Br₂
- Decomposition: Mg(ClO₃)₂ → MgCl₂ + 3 O₂
- 14) Synthesis: $2 H_2 + O_2 \rightarrow 2 H_2O$
- 15) Double replacement: 3 Ca(OH)₂ + 2 FeCl₃ → 2 Fe(OH)₃ + 3 CaCl₂
- 16) Single replacement: $Fe + Cu(NO_3)_2 \rightarrow Fe(NO_3)_2 + Cu$ [Assume Fe^{2+}]
- 17) Decomposition: 2 Hg₂O → 4 Hg + O₂
- 18) Synthesis: $S + O_2 \rightarrow SO_2$
- Double replacement: AgNO₃ + KI → AgI + KNO₃
- 20) Single replacement: Cu + 2 AgNO₃ → 2 Ag + Cu(NO₃)₂ [Copper (II) is used here]

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- 1. Understanding the eBook Crack8ing Chemistry Codes Worksheet Answer Key
 - The Rise of Digital Reading Crack8ing Chemistry Codes Worksheet Answer Key
 - o Advantages of eBooks Over Traditional Books
- 2. Identifying Crack8ing Chemistry Codes Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Crack8ing Chemistry Codes Worksheet Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Crack8ing Chemistry Codes Worksheet Answer Key

- Personalized Recommendations
- Crack8ing Chemistry Codes Worksheet Answer Key User Reviews and Ratings
- Crack8ing Chemistry Codes Worksheet Answer Key and Bestseller Lists
- 5. Accessing Crack8ing Chemistry Codes Worksheet Answer Key Free and Paid eBooks
 - Crack8ing Chemistry Codes Worksheet Answer Key Public Domain eBooks
 - Crack8ing Chemistry Codes Worksheet Answer Key eBook Subscription Services
 - Crack8ing Chemistry Codes Worksheet Answer Key Budget-Friendly Options
- 6. Navigating Crack8ing Chemistry Codes Worksheet Answer Key eBook Formats
 - o ePub, PDF, MOBI, and More
 - Crack8ing Chemistry Codes Worksheet Answer Key Compatibility with Devices
 - Crack8ing Chemistry Codes Worksheet Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Crack8ing Chemistry Codes Worksheet Answer Key
 - Highlighting and Note-Taking Crack8ing Chemistry Codes Worksheet Answer Key
 - Interactive Elements Crack8ing Chemistry Codes Worksheet Answer Key
- 8. Staying Engaged with Crack8ing Chemistry Codes Worksheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Crack8ing Chemistry Codes Worksheet Answer Key
- 9. Balancing eBooks and Physical Books Crack8ing Chemistry Codes Worksheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Crack8ing Chemistry Codes Worksheet Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Crack8ing Chemistry Codes Worksheet Answer Key
 - Setting Reading Goals Crack8ing Chemistry Codes Worksheet Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Crack8ing Chemistry Codes Worksheet Answer Key

- Fact-Checking eBook Content of Crack8ing Chemistry Codes Worksheet Answer Key
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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