

Exam #10 Review: Acids, Bases, and pH

1. Know the properties of acids and bases. What are their similarities and differences?

Acids

- Taste *sour*
- React with metals to form hydrogen gas
- Contain a hydrogen ion (H^{+1})
- pH value less than 7

Bases

- Taste *bitter*
- Feel slippery
- Contain a hydroxide ion (OH^{-1})
- pH value greater than 7

BOTH

- change colors of indicators
- react with each other to form salt and water
- conduct electricity when dissolved in solution (electrolytes)

2. Know the rules for naming
- acids**
- (binary acid and oxyacid) and name the following below:

Binary acids (only 2 elements): hydro – element – ic acidOxyacids (contain a polyatomic ion): NO "HYDRO" IN THE NAME

- -ate ion $\rightarrow$ -ic acid
- -ite ion $\rightarrow$ -ous acid

REMEMBER:

"I ate something icky,
all nite I was nauseous"

- a. HCl **hydrochloric acid**
b. HF **hydrofluoric acid**
c. $HClO_2$ **chlorous acid**

- d. HNO_3 **nitric acid**
e. H_3PO_4 **phosphoric acid**
f. H_2SO_3 **sulfurous acid**

3. Know the rules for naming
- bases**
- and name the following below:

Element + hydroxide

- a. NaOH **sodium hydroxide**
b. LiOH **lithium hydroxide**

- c. $Al(OH)_3$ **aluminum hydroxide**
d. $Ba(OH)_2$ **barium hydroxide**

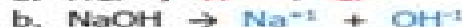
4. How are Arrhenius acids and bases defined?

Acids – produce H^{+1} ions in aqueous solutions**Bases** – produce OH^{-1} ions in aqueous solutions

5. How are Bronsted-Lowry acids and bases defined?

Acids – donate (give away) a H^{+1} ion (proton)**Bases** – accept (take in) a H^{+1} ion (proton)

6. Show how the following acids and bases ionize:



7. If it is a strong acid or base, how much do they ionize in solution?

STRONG acids & bases **COMPLETELY** ionize (all break apart) in a solution

8. If it is a weak acid or base, how much do they ionize in solution?

WEAK acids & bases do **NOT** completely ionize

9. What is a neutralization reaction? What products are always produced?

Acid reacts with a base**always produces water and a salt**

10. Complete the following neutralization reaction and what type of salt is made:

Type of Salt: neutralType of Salt: neutral

Acids And Bases Exam Review Guide Answers

Mark Dutton



Acids And Bases Exam Review Guide Answers:

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Table of Contents Acids And Bases Exam Review Guide Answers

1. Understanding the eBook Acids And Bases Exam Review Guide Answers
 - The Rise of Digital Reading Acids And Bases Exam Review Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Acids And Bases Exam Review Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acids And Bases Exam Review Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acids And Bases Exam Review Guide Answers
 - Personalized Recommendations
 - Acids And Bases Exam Review Guide Answers User Reviews and Ratings
 - Acids And Bases Exam Review Guide Answers and Bestseller Lists
5. Accessing Acids And Bases Exam Review Guide Answers Free and Paid eBooks
 - Acids And Bases Exam Review Guide Answers Public Domain eBooks
 - Acids And Bases Exam Review Guide Answers eBook Subscription Services

- Acids And Bases Exam Review Guide Answers Budget-Friendly Options
- 6. Navigating Acids And Bases Exam Review Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Acids And Bases Exam Review Guide Answers Compatibility with Devices
 - Acids And Bases Exam Review Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acids And Bases Exam Review Guide Answers
 - Highlighting and Note-Taking Acids And Bases Exam Review Guide Answers
 - Interactive Elements Acids And Bases Exam Review Guide Answers
- 8. Staying Engaged with Acids And Bases Exam Review Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acids And Bases Exam Review Guide Answers
- 9. Balancing eBooks and Physical Books Acids And Bases Exam Review Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acids And Bases Exam Review Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acids And Bases Exam Review Guide Answers
 - Setting Reading Goals Acids And Bases Exam Review Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acids And Bases Exam Review Guide Answers
 - Fact-Checking eBook Content of Acids And Bases Exam Review Guide Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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