Chapter 19 Study Guide: Acids, Bases & Titration - Answers

Differentiate between Arrhenius, Brønsted-Lowry & Lewis acids and bases.

	ACIDS	BASES	
Arrhenius	Form hydronium ions (H ₂ O*) in aqueous sol'n	Form hydroxide ions (OH·) in aqueous sol'n	
Brønsted-Lowry	Proton (H*) donors	Proton (H*) acceptors	
Lewis	Electron pair acceptors	Electron pair donors	

Label the acid (A), base (B), conjugate acid (CA) and conjugate base (CB) in each of the following reactions:

	Α	B	CB.	CA
21	H2CO3 +	CaCl ₂ →	CarCO ₃ +	2HCI
	A	В	CB	CA
Ib.	2 HCl +	Ca(OH)z	CaCla	+ 2 H ₂ O

Write the names or formulas for the following acids and bases:

a.	H ₂ S - hydrosulfuric acid
b.	HC2H3O2-acetic acid
C.	NaOH - sodium hydroxide
di	HNO: - nitric acid

- e. Nitrous acid HNO₂
- f. Calcium hydroxide Ca(OH)₂
- g hydrochloric acid HCl
- Use the following formulas for the remaining questions: (these will <u>not</u> be given to you on the test!)

$$K_w = [H_3O^*][OH^*] = 1.0 \times 10^{-14}$$

 $pH = -log[H_3O^*]$
 $pOH = -log[OH^*]$

$$pH + pOH = 14$$

$$M_A \cdot V_A \cdot n_A = M_B \cdot V_B \cdot n_B$$

4. What is the pH of 1.34 x 10⁻⁴ M solution of hydrochloric acid?

$$pH = -log [1.34 \times 10^{-4}M] = 3.873$$

Determine the pH and pOH of a 1.85 x 10⁻⁶ M NaOH solution.

$$pOH = -log [1.85 \times 10-6M] = 5.733$$

 $pH = 14 - 5.73 = 8.267$

What is the [OH·] of an aqueous solution that has a pH of 5.0?

7. What is the pH and the pOH of a solution with a volume of 2000. mL containing 35 grams of NaOH? Is this solution acidic or basic?

$$\frac{35gNaOH}{2.000L} \times \frac{1molNaOH}{40.00gNaOH} = 0.44M NaOH$$

pOH = $-\log[0.44M] = 0.36$ pH = $14 - 0.36 = 13.64$ basic

Acids And Bases React Review Guide Answers

Arshad Iqbal

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Table of Contents Acids And Bases React Review Guide Answers

- 1. Understanding the eBook Acids And Bases React Review Guide Answers
 - The Rise of Digital Reading Acids And Bases React Review Guide Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Acids And Bases React Review Guide Answers
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - $\circ \ \ Determining \ Your \ Reading \ Goals$
- 3. Choosing the Right eBook Platform
 - $\circ \ \ Popular \ eBook \ Platforms$
 - Features to Look for in an Acids And Bases React Review Guide Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Acids And Bases React Review Guide Answers
 - Personalized Recommendations

- Acids And Bases React Review Guide Answers User Reviews and Ratings
- Acids And Bases React Review Guide Answers and Bestseller Lists
- 5. Accessing Acids And Bases React Review Guide Answers Free and Paid eBooks
 - o Acids And Bases React Review Guide Answers Public Domain eBooks
 - Acids And Bases React Review Guide Answers eBook Subscription Services
 - Acids And Bases React Review Guide Answers Budget-Friendly Options
- 6. Navigating Acids And Bases React Review Guide Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Acids And Bases React Review Guide Answers Compatibility with Devices
 - Acids And Bases React Review Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acids And Bases React Review Guide Answers
 - Highlighting and Note-Taking Acids And Bases React Review Guide Answers
 - Interactive Elements Acids And Bases React Review Guide Answers
- 8. Staying Engaged with Acids And Bases React Review Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acids And Bases React Review Guide Answers
- 9. Balancing eBooks and Physical Books Acids And Bases React Review Guide Answers
 - $\circ\,$ Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acids And Bases React Review Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acids And Bases React Review Guide Answers
 - Setting Reading Goals Acids And Bases React Review Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acids And Bases React Review Guide Answers
 - Fact-Checking eBook Content of Acids And Bases React Review Guide Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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