

5

ELECTRONS IN ATOMS

SECTION 5.1 MODELS OF THE ATOM (pages 127–132)

This section summarizes the development of atomic theory. It also explains the significance of quantized energies of electrons as they relate to the quantum mechanical model of the atom.

► The Development of Atomic Models (pages 127–128)

1. Complete the table about atomic models and the scientists who developed them.

Scientist	Model of Atom
Dalton	The atom is a solid indivisible mass.
Thomson	The atom is a ball of positive charge with electrons stuck into the ball.
Rutherford	Most of an atom's mass is concentrated in the small, positively charged nucleus. The electrons surround the nucleus and the rest of the atom is empty space.
Bohr	Electrons are arranged in concentric circular paths around the nucleus.

2. Is the following sentence true or false? The electrons in an atom can exist between energy levels. false

► The Bohr Model (pages 128–129)

3. What is a small, discrete unit of energy called?
It is called a quantum.
4. Circle the letter of the term that completes the sentence correctly: A quantum of energy is the amount of energy required to
- move an electron from its present energy level to the next lower one
 - maintain an electron in its present energy level
 - c.** move an electron from its present energy level to the next higher one
5. In general, the higher the electron is on the energy ladder, the farther it is from the nucleus.

5 Electrons In Atoms Guided Answers

Clyde Metz



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Table of Contents 5 Electrons In Atoms Guided Answers

1. Understanding the eBook 5 Electrons In Atoms Guided Answers
 - The Rise of Digital Reading 5 Electrons In Atoms Guided Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 5 Electrons In Atoms Guided Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 5 Electrons In Atoms Guided Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 5 Electrons In Atoms Guided Answers
 - Personalized Recommendations
 - 5 Electrons In Atoms Guided Answers User Reviews and Ratings
 - 5 Electrons In Atoms Guided Answers and Bestseller Lists
5. Accessing 5 Electrons In Atoms Guided Answers Free and Paid eBooks
 - 5 Electrons In Atoms Guided Answers Public Domain eBooks
 - 5 Electrons In Atoms Guided Answers eBook Subscription Services

- 5 Electrons In Atoms Guided Answers Budget-Friendly Options
- 6. Navigating 5 Electrons In Atoms Guided Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 5 Electrons In Atoms Guided Answers Compatibility with Devices
 - 5 Electrons In Atoms Guided Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 5 Electrons In Atoms Guided Answers
 - Highlighting and Note-Taking 5 Electrons In Atoms Guided Answers
 - Interactive Elements 5 Electrons In Atoms Guided Answers
- 8. Staying Engaged with 5 Electrons In Atoms Guided Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 5 Electrons In Atoms Guided Answers
- 9. Balancing eBooks and Physical Books 5 Electrons In Atoms Guided Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 5 Electrons In Atoms Guided Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 5 Electrons In Atoms Guided Answers
 - Setting Reading Goals 5 Electrons In Atoms Guided Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 5 Electrons In Atoms Guided Answers
 - Fact-Checking eBook Content of 5 Electrons In Atoms Guided Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
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