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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23 $\text{Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$
oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$ reduction $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$ 24 3 $\text{Ag} + \text{Ni}^{3+} \rightarrow \text{Ni} + 3\text{Ag}^{+}$ oxidation reduction 25 $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 $\text{O}_2 + 2\text{H}_2 \rightarrow 2\text{H}_2\text{O}$

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ Sr^0 to Sr^2 oxidized reducing agent O^0 to O^2 reduced ox ag $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ Li^0 to Li^1 oxidized red ag S^0 to S^2 reduced ox ag $3\text{Cs} + \text{Br}_2$

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