Study Buddy: Atomic Structure and Nuclear Chemistry

Scientist	Experiment	Findings	Limitations	Model
Democritus	No experiment	Atoms are the smallest form of matter and cannot be created, destroyed or divided	Did not explain chemical behavior	
Dalton	Use experiments of previous scientists	Postulates: Atoms of elements are identical; atoms can't be created or destroyed; atoms combine in whole number ratios in a chemical reaction	Did not include sub- atomic particles	
Thomson	Cathode ray	Discovered the electron "Plum- pudding" Model	Did not explain deflection pattern of alpha particles as seen by Rutherford	
Rutherford	Gold foil	Discovered the proton and nucleus. Atoms contain mostly empty space.	Could not explain why electrons do not crash into the nucleus	0
Chadwick	Alpha particles	Discovered the neutron	No modification of atom model	No new model/neutron later add to nucleus
Bohr	Uses findings from previous scientists and Quantum theory	Electrons are in orbits of differing energy levels.	Does not explain the behavior of other elements just hydrogen	(3)
Modern Atom	Multiple experiments	Electron cloud and dual nature of electrons	Still doesn't explain the exceptions	· ===

Parts of an Atom (8th Grade Review)

Sub-atomic Particle	Location	Charge	Mass
Proton (p*)	Nucleus	1+	1
Neutron (n*)	Nucleus	0	1
Electron (er)	outside the nucleus	1-	0 or 1/1840

Nuclear Symbol Carbon-14

Protons:	- 6	
Neutrons:	-8	
Atomic Number:	6	
Mass Number:	14	

Identify the proton, electron, neutron in the model below. electron neutron proton What is the atomic number of the above models? 2 What is the atomic mass? 4.00 amu What element is this? Helium

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Table of Contents Atomic Structure Study Guide And Nuclear Processes

- 1. Understanding the eBook Atomic Structure Study Guide And Nuclear Processes
 - The Rise of Digital Reading Atomic Structure Study Guide And Nuclear Processes
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Atomic Structure Study Guide And Nuclear Processes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Atomic Structure Study Guide And Nuclear Processes
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Atomic Structure Study Guide And Nuclear Processes
 - Personalized Recommendations
 - Atomic Structure Study Guide And Nuclear Processes User Reviews and Ratings
 - Atomic Structure Study Guide And Nuclear Processes and Bestseller Lists
- 5. Accessing Atomic Structure Study Guide And Nuclear Processes Free and Paid eBooks
 - o Atomic Structure Study Guide And Nuclear Processes Public Domain eBooks
 - Atomic Structure Study Guide And Nuclear Processes eBook Subscription Services
 - Atomic Structure Study Guide And Nuclear Processes Budget-Friendly Options

- 6. Navigating Atomic Structure Study Guide And Nuclear Processes eBook Formats
 - o ePub, PDF, MOBI, and More
 - Atomic Structure Study Guide And Nuclear Processes Compatibility with Devices
 - Atomic Structure Study Guide And Nuclear Processes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Atomic Structure Study Guide And Nuclear Processes
 - Highlighting and Note-Taking Atomic Structure Study Guide And Nuclear Processes
 - Interactive Elements Atomic Structure Study Guide And Nuclear Processes
- 8. Staying Engaged with Atomic Structure Study Guide And Nuclear Processes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Atomic Structure Study Guide And Nuclear Processes
- 9. Balancing eBooks and Physical Books Atomic Structure Study Guide And Nuclear Processes
 - $\circ\,$ Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Atomic Structure Study Guide And Nuclear Processes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Atomic Structure Study Guide And Nuclear Processes
 - Setting Reading Goals Atomic Structure Study Guide And Nuclear Processes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Atomic Structure Study Guide And Nuclear Processes
 - Fact-Checking eBook Content of Atomic Structure Study Guide And Nuclear Processes
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

• Interactive and Gamified eBooks

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