

Bonding Test Study Guide

1. Define electrostatic attraction, crystal lattice, lattice energy and electrolyte.

2. What ions do the following elements form? (Examples: Al forms Al^{3+} and Cl forms Cl^{-})

- a. O b. Ca c. Na d. F e. N

3. Ionic bonds are formed when electrons are _____ from the _____ to the _____. How does an ionic bond form? (use crystal lattice, electrostatic attraction and transfer in your answer)

4. Do metals tend to form cations or anions? How about nonmetals?

5. What are some properties of ionic compounds? Covalent compounds?

6. Covalent bonds are formed between _____ that _____ their electrons.

7. Fill in the following chart:

	Lewis Dot Structure	EN difference	Shape	Bond Polarity	Molecular polarity
a. CH_4					
b. OH^{-}					
c. H_2O					
d. CO_2					
e. NH_3					
f. SO_2					
g. NO_3^{-}					

8. Draw the Lewis structures for the following:

- a. strontium b. dihydrogen selenide c. carbon disulfide
d. carbon monoxide e. chloride

9. What causes metals to have the properties that they have (what do they do with their electrons)? How can you describe the valence electrons in metals?

10. How do electronegativities relate to the type of bond that is formed between two elements?

11. Define nonpolar covalent, polar covalent and dipole.

12. Predict the formulas for the following compounds (transfer these valence electrons):

- a. sodium chloride b. diphosphorus pentoxide c. carbon dioxide d. aluminum oxide

13. Name the following compounds:

- a. KBr b. H_2O c. Ca_3O d. Na_3N e. C_6H_6

14. When bonds break, energy is _____. When bonds are made, energy is _____.

15. Study your polyatomic ions!!!

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