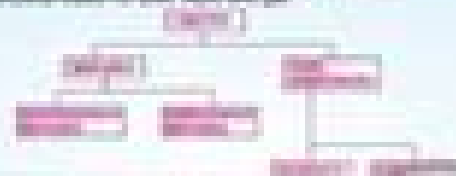


1. CLASSIFICATION OF MATTER

- Matter can exist in three physical states: solid, liquid and gas.



- Various properties of matter are used to classify matter into these categories.
- Matter is classified into pure substance and mixture.
- Pure substance is further classified into element and compound.
- Mixture is further classified into homogeneous mixture and heterogeneous mixture.
- Homogeneous mixture is further classified into solution and alloy.
- Heterogeneous mixture is further classified into suspension and colloid.

2. PROPERTIES OF MATTER AND THEIR MEASUREMENT

- Properties of matter are classified into physical and chemical properties.
 - Physical properties are those which can be observed or measured without changing the identity of the substance.
 - Chemical properties are those which can be observed or measured by changing the identity of the substance.
 - Colour, odour, melting point, density, etc. are some of the physical properties.
 - Combustion, corrosion, etc. are some of the chemical properties.
 - Quantitative measurement of physical properties is necessary for scientific study.
 - Properties of matter are measured in SI units.
 - Some important quantities and their units are:
- | Quantity | SI Unit |
|------------------|------------------------------------|
| Mass (g) | Kilogram (kg) |
| Volume (mL) | Cubic decimetre (dm ³) |
| Temperature (°C) | Kelvin (K) |

Some important units and their conversions

1 kg = 1000 g	1 g = 1000 mg	1 kg = 1000 g	1 kg = 1000 g	1 kg = 1000 g
1 kg = 1000 g	1 g = 1000 mg	1 kg = 1000 g	1 kg = 1000 g	1 kg = 1000 g

Some important units

- 1 kg = 1000 g, 1 g = 1000 mg, 1 kg = 1000 g, 1 kg = 1000 g, 1 kg = 1000 g

3. UNCERTAINTY IN MEASUREMENT

- Measurement is a comparison of a quantity with a standard unit.
- Measurement is always uncertain.
- Measurement is always uncertain.

- Significant figures are those digits which are known with certainty plus one additional uncertain or doubtful digit.

Rules for determining the number of significant figures

- Non-zero digits are significant.
- Zeros between two non-zero digits are significant.
- Zeros at the end or right of a number are significant provided they are on the right side of a decimal point.
- Leading zeros are not significant figures.
- In a number written in scientific notation, all digits are significant.
- Exact values are the numbers of atoms, molecules, etc. in a sample.
- Round-off to the nearest integer of a positive value is the first value of the next digit.
- In addition and subtraction of significant figures, the result cannot have more digits to the right of the decimal point than either of the original numbers.
- In multiplication and division of significant figures, the result must be rounded off to the same number of significant figures as the least number of significant figures in the original numbers.

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RJ Shavelson



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