

Quiz & Worksheet - Faraday's Law of Electrolysis



- o mis proportional to Z
- m is proportional to 1/I.
- m is proportional to 1/E
- m is proportional to 1/t

In an electrolysis, a current of 5.6A passes through an electrolyte (potassium sulfate) for 3 hours. Calculate the mass of potassium produced.

- © 245g
- m 39.8g
- a 32.8g
- o 6.7g

3. What amount of electricity must be supplied to an electrolytic cell in order to produce 0.100Kg of Iron from Iron(III)oxide?

- m 96500C
- m 4260C
- o 515962C
- m 804550C

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Faraday S Law Worksheet With Answers

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