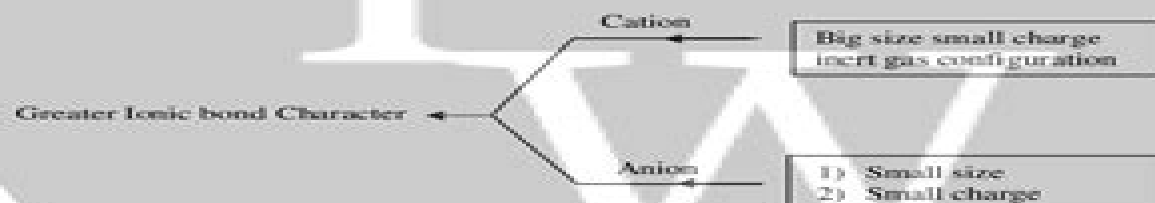


CHAPTER

04

CHEMICAL BONDING & MOLECULAR STRUCTURE

- An atom (or) group of atoms having independent existence in nature is defined as molecule.
- Force of attraction that holds two atoms (or) oppositely charged ions in molecule is chemical bond.
- An atom must have eight electrons in its outermost energy level for its stability is called as octet rule.
- The number of electrons an atom loses (or) gains in an ionic bond formation is called electrovalency.
- According to Fajan's Rules
 - The no. of electrons an atom of the element loses (or) gains is known as electrovalency of that element.
 - Favourable conditions for cation formation.
 - 1) Low I.P
 - 2) Low charge on ion
 - 3) Big atomic size
 - 4) Inert gas configuration
 - Favourable conditions for anion formation.
 - 1) High E.N & E.A
 - 2) Small atomic size
 - 3) Low charge on ion



- LiCl is more covalent in nature and so dissolved in non-polar solvents like alcohol (or) ether.
- Ions which have 18 electrons $[ns^2 np^6 nd^{10}]$ in their outermost main energy level are stable and it is called as pseudo inert gas configuration (or) nickel group configuration.
- The energy liberated when an ionic crystal is formed from an assembly of isolated gaseous ions is called lattice energy.
- Lattice energy

$$(\cup) = \left[-N_0 \frac{AZ^+ Z^- e^2}{r} + N_0 \frac{Be^2}{r^9} \right]$$
 - $$\cup = \frac{AZ^+ Z^- e^2 N_0}{\epsilon_0} \left[1 - \frac{1}{n} \right]$$

(Born – Lande equation)

Coordination number of Na^+ and Cl^- in NaCl crystal is (6:6)

Coordination number of Cs^+ and Cl^- in NaCl crystal is (8:8)
 - The total amount of energy change in a reaction remains the same whether the reaction take place in one (or) several stages is known as Hess law of constant heat summation.

Bonding Chemical Formulas

John Olmsted, Gregory M. Williams



Bonding Chemical Formulas:

Fundamentals of Environmental Chemistry, Third Edition Stanley E. Manahan, 2011-03-05 Written by an expert using the same approach that made the previous two editions so successful *Fundamentals of Environmental Chemistry Third Edition* expands the scope of book to include the strongly emerging areas broadly described as sustainability science and technology including green chemistry and industrial ecology The new edition includes Increased emphasis on the applied aspects of environmental chemistry Hot topics such as global warming and biomass energy Integration of green chemistry and sustainability concepts throughout the text More and updated questions and answers including some that require Internet research Lecturers Pack on CD ROM with solutions manual PowerPoint presentations and chapter figures available upon qualifying course adoptions The book provides a basic course in chemical science including the fundamentals of organic chemistry and biochemistry The author uses real life examples from environmental chemistry green chemistry and related areas while maintaining brevity and simplicity in his explanation of concepts Building on this foundation the book covers environmental chemistry broadly defined to include sustainability aspects green chemistry industrial ecology and related areas These chapters are organized around the five environmental spheres the hydrosphere atmosphere geosphere biosphere and the anthrosphere The last two chapters discuss analytical chemistry and its relevance to environmental chemistry Manahan's clear concise and readable style makes the information accessible regardless of the readers level of chemistry knowledge He demystifies the material for those who need the basics of chemical science for their trade profession or study curriculum as well as for readers who want to have an understanding of the fundamentals of sustainable chemistry in its crucial role in maintaining a livable planet

Fundamentals of Sustainable Chemical Science Stanley E. Manahan, 2009-03-10 Written by Stanley Manahan *Fundamentals of Sustainable Chemical Science* has been carefully designed to provide a basic introduction to chemistry including organic chemistry and biochemistry for readers with little or no prior background in the subject Manahan bestselling author of many environmental texts presents the material in a practical

Chemistry John Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science

Chemical Bonding in Crystals and Their Properties Anna V. Vologzhanina, Yulia V. Nelyubina, 2020-06-16 Unravelling an intricate network of interatomic interactions and their relations to different behaviors of chemical compounds is key to the successful design of new materials for both existing and novel applications from medicine to innovative concepts of molecular electronics and spintronics X ray crystallography has proven to be very helpful in addressing many important chemical problems in modern materials science and biosciences Intertwined with computational techniques it provides insights into the nature of chemical bonding and the physicochemical properties including optical magnetic electrical mechanical and others of crystalline materials otherwise accessible by experimental techniques that are not so readily available to chemists In addition to the advanced approaches in charge density analysis made possible by X ray diffraction the information

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Forensic Chemistry JaVed I. Khan, Thomas J. Kennedy, Donnell R. Christian, Jr., 2011-11-15 This book focuses on a marvel approach that blends chemistry with forensic science and is used for the examination of controlled substances and clandestine operations The book will particularly interest forensic chemists forensic scientists criminologists and biochemists

Orbitals in Chemistry Victor M. S. Gil, 2000-08-10 This text presents a unified and up to date discussion of the role of atomic and molecular orbitals in chemistry from the quantum mechanical foundations to the recent developments and applications The discussion is mainly qualitative largely based on symmetry arguments It is felt that a sound mastering of the concepts and qualitative interpretations is needed especially when students are becoming more and more familiar with numerical calculations based on atomic and molecular orbitals The text is mathematically less demanding than most traditional quantum chemistry books but still retains clarity and rigour The physical insight is maximized and abundant illustrations are used The relationships between the more formal quantum mechanical formalisms and the traditional chemical descriptions of chemical bonding are critically established This book is of primary interest to undergraduate chemistry students and others taking courses of which chemistry is a significant part

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Anne Tindale, 1998 NO description available

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Southam, Gwendolyn Lawrie, Natalie Williamson, Christopher Thompson, 2024-01-09 The third edition of Chemistry Core

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Zhen-Yu Tian, 2025-03-08 This book is a comprehensive guide that explores the intricate relationship between chemistry and the environment The book offers a systematic study of various aspects including basic chemical concepts atmospheric chemistry phase interactions water pollution and treatment the atmosphere and its chemistry particles in the atmosphere photochemical smog the geosphere geochemistry and soil chemistry It offers a blend of theoretical explanations practical examples detailed illustrations and tables to facilitate a deeper understanding of the subject matters The book's uniqueness is its ability to simplify complex scientific concepts and present them in an accessible and engaging manner It employs a didactic approach making learning a proactive and enjoyable process Readers appreciate the in depth exploration of topics such as the law of conservation of matter the role of the atmosphere in safeguarding life on Earth the formation of photochemical smog the evolution of the geosphere and the role of soil colloids in soil chemistry It emphasizes the importance of understanding these concepts to address environmental challenges effectively The key benefit for readers is gaining a comprehensive understanding of environmental chemistry its concepts and its real world applications This book serves as an essential resource for those aiming to gain in depth understanding of the subject and its significance in preserving our planet This makes it particularly appealing to students researchers and professionals interested in environmental sciences **Chemistry** John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk

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