

## 2 - Bonding & Molecular Structures

10 20 30 40 50 60 70 80 90 100

Table 4.4 - Some Average Single and Multiple Bond Strengths (kJ/mol)

	H	C	N	O	F	Cl	Br	I	S	Se	Te
H	436										
C		413	338	358	479						
N			391	358	480						
O				498	500						
F					485						
Cl						328					
Br							276				
I								240			
S									265		
Se										226	
Te											226

### Multiple Bonds

C=C	614	C≡C	839
C=O	799	C≡N	891
C≡N	891	N≡N	945
O=O	498		
O=O and N=O	500		

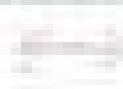
Table 4.5 - Standard Enthalpies of Formation (kJ/mol)

Compound	Standard Enthalpy of Formation	Compound	Standard Enthalpy of Formation
H <sub>2</sub> O(l)	-285.8	H <sub>2</sub> O(g)	-241.8
H <sub>2</sub> O(s)	-291.8		

1. Write the balanced chemical equation for the complete combustion of ethane (C<sub>2</sub>H<sub>6</sub>).



2. Draw structural formula for ethane, showing all bonds and lone pairs.



3. Calculate the standard enthalpy of formation for the products in the reaction:  
 Calculate the standard enthalpy of formation for the products in the reaction.



4. What is the standard enthalpy of formation for CO<sub>2</sub> (g)?

5. Use the bond enthalpies provided to calculate the standard enthalpy of formation for CO<sub>2</sub> (g).

# Bonding And Molecular Structure Worksheet Answers South Pasadena 9

**George E. Ryschkewitsch**



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