

Student Exploration: Bohr Model of Hydrogen

Vocabulary: absorption spectrum, Bohr model, electron volt, emission spectrum, energy level, ionization energy, laser, orbital, photon

Prior Knowledge Questions (Answer these questions BEFORE using the Gizmo.)

1. What happens to an electron when it absorbs a tiny packet of light, called a photon?

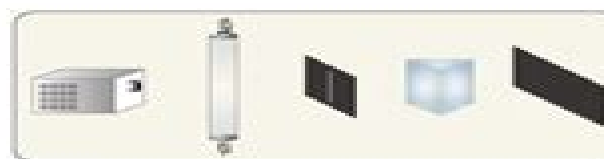
It will move up to a higher energy level.

2. What happens to an electron when it emits a photon?

It will move down to a lower energy level.

Gizmo Warm-up

When white light is passed through hydrogen gas and separated by a prism, some wavelengths of light are absorbed by the hydrogen atoms in the gas. This causes black bars to appear in the **absorption spectrum** of hydrogen. An **emission spectrum** is created when hydrogen gas emits light.



The **Bohr model** of the atom was inspired by the spectrum produced by hydrogen gas. The link between light and atomic structure is illustrated by the *Bohr Model of Hydrogen* Gizmo. The Gizmo shows a **laser** pointed at a container of hydrogen gas. After passing through the gas, light from the laser goes through a prism and is detected on a screen.

1. With the **Laser energy** set to 7.0 eV, click **Play** () . Observe the red electron on the **Orbitals** pane. **Do 7.0-eV photons have any effect on the electron?** Photons do not have any effect on the electron as it stays at energy level 1.
2. Click **Pause** () . Set the **Laser energy** to 12.1 eV and click **Play**. **What happens to the electron now?** The electron has moved up to energy level 3.
3. Under **Go to energy level**, select **1**. Notice that a photon is emitted by the electron. **What is the energy of the emitted photon?** 12.1 eV

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