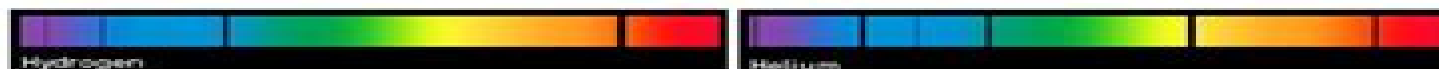


## Student Exploration: Bohr Model: Introduction

When light passes through a gas, certain wavelengths of the light are absorbed. The result is a unique **absorption spectrum**. Two examples are shown below.



1. What colors of light are absorbed by hydrogen gas?  
violet, blue, green, and red
2. What colors of light are absorbed by helium gas?  
violet, blue, green, yellow, and red

### Gizmo Warm-up

In 1913, Niels Bohr proposed that the unique spectral lines created by different elements were related to the way electrons were arranged around the nucleus. The *Bohr Model: Introduction* Gizmo explores this connection.

The **laser** shown in the Gizmo can emit **photons**, or particles of light, at a variety of wavelengths. The energy of a photon, measured in **electron volts (eV)**, is inversely proportional to its wavelength. Photons that pass through the gas are detected by the photon detector at the right.



1. With the **Energy (eV)** set to 1 eV, click **Fire**. Did the photon go straight through the gas in the tube, or was it absorbed by the gas?  
the photon went straight through the gas tube.
2. Set the **Energy (eV)** to 4 eV, and click **Fire**. What happened this time?  
the photon was absorbed by the gas.

### Activity A: Absorption spectra

#### Get the Gizmo ready:

- On the **SIMULATION** pane, select **Lamp**.
- Check that **Gas C** is selected.



**Introduction:** The smaller the wavelength of a photon, the greater its energy. Humans can see photons with wavelengths between 700 nanometers (red) and 400 nanometers (violet), corresponding to energies between 1.8 and 3.1 eV.

**Question:** What does the absorption spectrum of an element indicate about its electron configuration?

# Bohr Gizmo Answers

**L Darling-Hammond**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the text 'L Darling-Hammond'.

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## **Table of Contents Bohr Gizmo Answers**

1. Understanding the eBook Bohr Gizmo Answers
  - The Rise of Digital Reading Bohr Gizmo Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Bohr Gizmo Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Bohr Gizmo Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Bohr Gizmo Answers
  - Personalized Recommendations
  - Bohr Gizmo Answers User Reviews and Ratings
  - Bohr Gizmo Answers and Bestseller Lists
5. Accessing Bohr Gizmo Answers Free and Paid eBooks
  - Bohr Gizmo Answers Public Domain eBooks
  - Bohr Gizmo Answers eBook Subscription Services
  - Bohr Gizmo Answers Budget-Friendly Options
6. Navigating Bohr Gizmo Answers eBook Formats

- ePub, PDF, MOBI, and More
- Bohr Gizmo Answers Compatibility with Devices
- Bohr Gizmo Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Bohr Gizmo Answers
  - Highlighting and Note-Taking Bohr Gizmo Answers
  - Interactive Elements Bohr Gizmo Answers
- 8. Staying Engaged with Bohr Gizmo Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Bohr Gizmo Answers
- 9. Balancing eBooks and Physical Books Bohr Gizmo Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Bohr Gizmo Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Bohr Gizmo Answers
  - Setting Reading Goals Bohr Gizmo Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Bohr Gizmo Answers
  - Fact-Checking eBook Content of Bohr Gizmo Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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