

QUESTION 7 (Start on a new page.)

Hydrated potassium carbonate, $K_2CO_3 \cdot xH_2O$, is a WEAK BASE. A solution is prepared by dissolving some of this solid in water.

7.1 Define the term weak base.

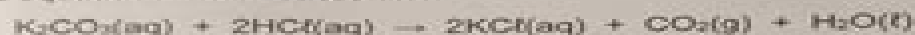
(2)

7.2 Write down the formula of the conjugate acid of the carbonate ion, $CO_3^{2-}(aq)$.

(1)

A hydrochloric acid solution, $HCl(aq)$, of concentration $0,1 \text{ mol} \cdot \text{dm}^{-3}$ is titrated with the prepared potassium carbonate solution, $K_2CO_3(aq)$, of unknown concentration.

The balanced equation for the reaction is:



The results of the titration are given below.

	VOLUME OF $HCl(aq)$ USED (cm^3)	$K_2CO_3(aq)$ IN BURETTE		VOLUME OF $K_2CO_3(aq)$ USED (cm^3)
		INITIAL BURETTE READING (cm^3)	FINAL BURETTE READING (cm^3)	
Run 1	25	6,5	p	20,05
Run 2	25	q	48,3	20,15

7.3 Determine the value of:

7.3.1 p

(1)

7.3.2 q

(1)

7.4 METHYL ORANGE is used as the indicator. Explain why methyl orange is the most suitable indicator for this titration by referring to the pH at the equivalence point.

(2)

7.5 Calculate the concentration of the K_2CO_3 solution.

(5)

The above K_2CO_3 solution used in the titration, was prepared by completely dissolving 6,525 g of the hydrated potassium carbonate, $K_2CO_3 \cdot xH_2O$, in 600 cm^3 water.

7.6 Calculate the value of x in the formula $K_2CO_3 \cdot xH_2O$.

(5)
[17]

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