



CHEMISTRY SSC-I

Revised Syllabus

Time allowed: 2:40 Hours

Total Marks Sections B and C: 53

NOTE:- Answer any eleven parts from Section 'B' and any two questions from Section 'C' on the separately provided answer book. Use supplementary answer sheet i.e. Sheet-B if required. Write your answers neatly and legibly.

SECTION – B (Marks 33)

Q. 2 Answer any ELEVEN parts. The answer to each part should not exceed three to four lines. (11 x 3 = 33)

- (i) Identify the branch of Chemistry, that is related to the following information:
 - a. Sulphuric Acid is weaker than HCl.
 - b. Gases can be compressed by applying pressure.
 - c. Hair contain a special class proteins called keratins which is present in nails and wool.
- (ii) Determine the formula masses of the following:
 - a. Urea $(\text{NH}_2)_2\text{CO}$
 - b. Ammonium Nitrate NH_4NO_3
- (iii) What were the conclusions of Rutherford's experiment?
- (iv) Write any three uses of Isotopes?
- (v) Identify the **group** and **period** of $^{27}_{13}\text{Al}$, ^9_5B , $^{24}_{12}\text{Mg}$ on the basis of electronic configuration.
- (vi) What is Atomic size? Write its trends in periodic table.
- (vii) What is the importance of Glue and Adhesives in our society?
- (viii) Differentiate between Diffusion and Effusion of Gases. Also give examples.
- (ix) Why does evaporation lower the temperature of Liquids?
- (x) Define **Saturated**, **Unsaturated** and **Super-saturated** solutions. Also give examples.
- (xi) Differentiate between Colloid and Suspension, with the help of examples.
- (xii) Define Oxidation Reduction reaction. Also give one example from daily life.
- (xiii) Draw a labelled diagram of the Daniel's cell.
- (xiv)
 - a. Which of the given elements is more metallic and Why? Mg or Al.
 - b. What is the importance of Mg?
 - c. F_2 oxidizes all the halide ions to free Halogens. How?
- (xv) Compare reactivity of Alkali and Alkaline earth metals by their positions in the periodic table.

SECTION – C (Marks 20)

Note: Attempt any TWO questions. All questions carry equal marks.

(2 x 10 = 20)

- Q. 3**
 - a. Define Electro-negativity. Explain the Electronegative characters of elements. Give examples. (04)
 - b. Give any three important uses of Platinum. (03)
 - c. What is Alloying? Also give example. (02)
 - d. Define Rusting of Iron. (01)
- Q. 4**
 - a. How can you manufacture Sodium Hydroxide from Brine? (04)
 - b. A solution is prepared by dissolving 10 gram of Haemoglobin in water to make up 1dm^3 in volume. Calculate molarity of this solution. Molar mass of Haemoglobin is $6.51 \times 10^4 \text{ g / mole}$. (04)
 - c. What is Concentration Unit? Give examples. (02)
- Q. 5**
 - a. What is Vapour Pressure? Explain effects of temperature on Vapour Pressure. (04)
 - b. Define Intermolecular forces. Explain Hydrogen Bonding found in Water. Also draw diagram. (04)
 - c. What is Shielding effect? Which atom has greater shielding effect and Why? Be or Mg. (02)

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