



CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules by sharing or transferring electrons. This process is called bonding. The resulting molecule is more stable than the individual atoms.

Types of Chemical Bonds:

- 1. Ionic Bond:** Formed by the transfer of electrons from one atom to another, resulting in oppositely charged ions that attract each other.
- 2. Covalent Bond:** Formed by the sharing of electron pairs between two atoms.
- 3. Metallic Bond:** Formed by the attraction between metal cations and a sea of delocalized electrons.

COVALENT BONDING AND MOLECULAR STRUCTURE

Covalent bonding occurs when two atoms share one or more pairs of electrons. This sharing results in a stable electron configuration for both atoms, often resembling the nearest noble gas.

Covalent bonds can be single, double, or triple bonds, depending on the number of electron pairs shared. Single bonds involve one pair of electrons, double bonds involve two pairs, and triple bonds involve three pairs.

The resulting molecule has a specific shape, determined by the arrangement of atoms and the repulsion between electron pairs. This shape affects the molecule's properties, such as its boiling point and reactivity.

SCIENCE REAL

A simplified periodic table with elements color-coded by groups: metals (shades of blue and green), nonmetals (shades of red and orange), and noble gases (yellow). It is used to illustrate trends in chemical bonding.

IONIC BONDING

Ionic bonding occurs when one atom transfers an electron to another atom, resulting in the formation of oppositely charged ions. These ions are then attracted to each other, forming an ionic bond.

COVALENT BONDING

Covalent bonding occurs when two atoms share one or more pairs of electrons. This sharing results in a stable electron configuration for both atoms, often resembling the nearest noble gas.

MOLECULAR STRUCTURE

The shape of a molecule is determined by the arrangement of atoms and the repulsion between electron pairs. This shape affects the molecule's properties, such as its boiling point and reactivity.

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Diagram illustrating the formation of a covalent bond:

Two atoms, each with one valence electron, combine to form a molecule. The two valence electrons are shared, forming a covalent bond. The resulting molecule has a stable electron configuration.

Diagram illustrating the formation of an ionic bond:

One atom with one valence electron transfers its electron to another atom with one valence electron. This results in the formation of oppositely charged ions, which are then attracted to each other, forming an ionic bond.

Diagram illustrating the formation of a metallic bond:

Many metal atoms combine to form a metal crystal. The valence electrons are delocalized, forming a sea of electrons that surrounds the metal cations. This results in the formation of a metallic bond.

GUIDED READING ACTIVITY

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Yi-Tong Ma



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1. Understanding the eBook Covalent Bonds Reading Notetaking Guide
 - The Rise of Digital Reading Covalent Bonds Reading Notetaking Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonds Reading Notetaking Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Bonds Reading Notetaking Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonds Reading Notetaking Guide
 - Personalized Recommendations
 - Covalent Bonds Reading Notetaking Guide User Reviews and Ratings
 - Covalent Bonds Reading Notetaking Guide and Bestseller Lists
5. Accessing Covalent Bonds Reading Notetaking Guide Free and Paid eBooks

- Covalent Bonds Reading Notetaking Guide Public Domain eBooks
 - Covalent Bonds Reading Notetaking Guide eBook Subscription Services
 - Covalent Bonds Reading Notetaking Guide Budget-Friendly Options
6. Navigating Covalent Bonds Reading Notetaking Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Covalent Bonds Reading Notetaking Guide Compatibility with Devices
 - Covalent Bonds Reading Notetaking Guide Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Bonds Reading Notetaking Guide
 - Highlighting and Note-Taking Covalent Bonds Reading Notetaking Guide
 - Interactive Elements Covalent Bonds Reading Notetaking Guide
 8. Staying Engaged with Covalent Bonds Reading Notetaking Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Bonds Reading Notetaking Guide
 9. Balancing eBooks and Physical Books Covalent Bonds Reading Notetaking Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonds Reading Notetaking Guide
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Covalent Bonds Reading Notetaking Guide
 - Setting Reading Goals Covalent Bonds Reading Notetaking Guide
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Covalent Bonds Reading Notetaking Guide
 - Fact-Checking eBook Content of Covalent Bonds Reading Notetaking Guide
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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