

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

|                                             |                         |
|---------------------------------------------|-------------------------|
| _____ <b>lattice energy</b> _____           | <b>a.</b> ionic bond    |
| _____ <b>bond energy</b> _____              | <b>b.</b> covalent bond |
| _____ <b>enthalpy of vaporization</b> _____ | <b>c.</b> metallic bond |

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

|                                   |                 |                                      |               |
|-----------------------------------|-----------------|--------------------------------------|---------------|
| _____ <b>ionic</b> _____          | <b>a.</b> H—F   | _____ <b>nonpolar covalent</b> _____ | <b>d.</b> H—H |
| _____ <b>ionic</b> _____          | <b>b.</b> Na—Cl | _____ <b>polar covalent</b> _____    | <b>e.</b> H—C |
| _____ <b>polar covalent</b> _____ | <b>c.</b> H—O   | _____ <b>polar covalent</b> _____    | <b>f.</b> H—N |

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen  
atom and the unshared electron pair of a strongly electronegative atom such as O,  
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,  
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally  
intermolecular, while ionic and covalent bonds occur between ions or atoms  
respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.  
Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making  
water a liquid at room temperature. Hydrogen bonding exists between water  
molecules, but not between hydrogen sulfide molecules.

# Covalent Bonding Worksheet Answers Chapter 8 Pg 19

**A Loxley**



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## Decoding **Covalent Bonding Worksheet Answers Chapter 8 Pg 19**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Covalent Bonding Worksheet Answers Chapter 8 Pg 19**," a mesmerizing literary creation penned with a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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