

Covalent Bonding Review Sheet – Dec 12, 2012

Name _____ class period _____

- Given the formula, a compound can be named using conventional systems that include Greek prefixes and Roman numeral where appropriate.
- Differences in electronegativity values can be used to predict where a bond fits on the continuum between ionic and covalent bonds.
- Since most compounds contain multiple bonds, a substance may contain more than one type of bond.

Section 1: Naming and Formula Writing

Write the formulas for the following covalent compounds

- | | |
|--------------------------|-----------------------------------|
| 1. Nitrogen Monoxide | <u>NO</u> |
| 2. Dinitrogen Monoxide | <u>N₂O</u> |
| 3. Nitrogen Dioxide | <u>NO₂</u> |
| 4. Dinitrogen Trioxide | <u>N₂O₃</u> |
| 5. Dinitrogen Tetraoxide | <u>N₂O₄</u> |
| 6. Dinitrogen Pentaoxide | <u>N₂O₅</u> |
| 7. carbon dioxide | <u>CO₂</u> |
| 8. phosphorus triiodide | <u>PI₃</u> |
| 9. sulfur dichloride | <u>SCl₂</u> |
| 10. nitrogen trifluoride | <u>NF₃</u> |

Write names from the following formulas

- | | |
|------------------------------------|------------------------------------|
| 11. N ₂ F ₄ | <u>dinitrogen tetrafluoride</u> |
| 12. SCl ₄ | <u>sulfur tetrachloride</u> |
| 13. ClF ₃ | <u>chlorine trifluoride</u> |
| 14. SiO ₂ | <u>silicon dioxide</u> |
| 15. P ₄ O ₁₀ | <u>tetraphosphorus decoxide</u> |
| 16. SiF ₄ | <u>silicon tetrafluoride</u> |
| 17. H ₂ O | <u>dihydrogen monoxide</u> |
| 18. Br ₂ | <u>dibromide</u> or <u>bromine</u> |
| 19. HBr | <u>hydrogen monobromide</u> |
| 20. N ₂ S ₃ | <u>dinitrogen trisulfide</u> |

Covalent Bonding Vocabulary Review Answer Key

SJ Ball



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Table of Contents Covalent Bonding Vocabulary Review Answer Key

1. Understanding the eBook Covalent Bonding Vocabulary Review Answer Key
 - The Rise of Digital Reading Covalent Bonding Vocabulary Review Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonding Vocabulary Review Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Bonding Vocabulary Review Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonding Vocabulary Review Answer Key
 - Personalized Recommendations
 - Covalent Bonding Vocabulary Review Answer Key User Reviews and Ratings

- Covalent Bonding Vocabulary Review Answer Key and Bestseller Lists
- 5. Accessing Covalent Bonding Vocabulary Review Answer Key Free and Paid eBooks
 - Covalent Bonding Vocabulary Review Answer Key Public Domain eBooks
 - Covalent Bonding Vocabulary Review Answer Key eBook Subscription Services
 - Covalent Bonding Vocabulary Review Answer Key Budget-Friendly Options
- 6. Navigating Covalent Bonding Vocabulary Review Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Covalent Bonding Vocabulary Review Answer Key Compatibility with Devices
 - Covalent Bonding Vocabulary Review Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Bonding Vocabulary Review Answer Key
 - Highlighting and Note-Taking Covalent Bonding Vocabulary Review Answer Key
 - Interactive Elements Covalent Bonding Vocabulary Review Answer Key
- 8. Staying Engaged with Covalent Bonding Vocabulary Review Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Bonding Vocabulary Review Answer Key
- 9. Balancing eBooks and Physical Books Covalent Bonding Vocabulary Review Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonding Vocabulary Review Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Covalent Bonding Vocabulary Review Answer Key
 - Setting Reading Goals Covalent Bonding Vocabulary Review Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Covalent Bonding Vocabulary Review Answer Key
 - Fact-Checking eBook Content of Covalent Bonding Vocabulary Review Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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