



COVALENT BONDING

Practice Problems

Do your best to do better than your previous problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Draw a molecular formula for each of the following compounds.

- a. Silicon tetrachloride b. Nitrogen dioxide c. Hydrogen peroxide
d. Sulfur hexafluoride e. Dinitrogen pentoxide

2. Write the formula for the following molecular compounds.

- a. Carbon monoxide b. Diphosphorus pentoxide c. Nitrogen
d. Sulfur e. Hydrogen f. Oxygen

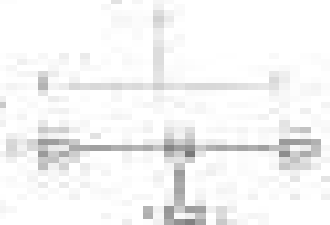
3. Draw a Lewis structure for each of the following molecular compounds.
 a. Diphosphorus pentoxide b. Carbon monoxide c. Hydrogen peroxide d. Sulfur hexafluoride
 e. Nitrogen dioxide f. Nitrogen monoxide g. Nitrogen trioxide h. Nitrogen pentoxide
 i. Nitrogen tetroxide j. Nitrogen trioxide k. Nitrogen dioxide l. Nitrogen monoxide
 m. Nitrogen pentoxide n. Nitrogen trioxide o. Nitrogen dioxide p. Nitrogen monoxide

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw a Lewis structure for each of the following compounds.



2. Draw a Lewis structure for each of the following compounds.



3. Draw a Lewis structure for each of the following compounds.



4. Draw a Lewis structure for each of the following compounds.

5. Draw a Lewis structure for each of the following compounds. Show all lone pairs and indicate the formal charge on each atom.



6. Draw a Lewis structure for each of the following compounds.

SECTION 8.3 BONDING: THE ORBITAL

1. Draw a Lewis structure for each of the following compounds. Show all lone pairs and indicate the formal charge on each atom.
 a. Nitrogen dioxide b. Nitrogen monoxide c. Nitrogen trioxide d. Nitrogen pentoxide
 e. Nitrogen tetroxide f. Nitrogen trioxide g. Nitrogen dioxide h. Nitrogen monoxide
 i. Nitrogen pentoxide j. Nitrogen trioxide k. Nitrogen dioxide l. Nitrogen monoxide
 m. Nitrogen pentoxide n. Nitrogen trioxide o. Nitrogen dioxide p. Nitrogen monoxide

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