

## Chemistry Test Answer 24

☐ A  
☐ B  
☒ C  
☐ D

The diagram shows a chemical reaction between two molecules. The first molecule is a hydrocarbon with a central carbon atom bonded to four other carbon atoms in a branched structure. The second molecule is a diatomic gas, represented by two identical atoms bonded together. The reaction arrow points to the right, indicating the reaction proceeds in that direction.

The reaction is exothermic, as indicated by the negative sign in the enthalpy change ( $\Delta H$ ). The reaction is also spontaneous, as indicated by the negative sign in the Gibbs free energy change ( $\Delta G$ ).

The reaction is a combustion reaction, as it involves the reaction of a hydrocarbon with a diatomic gas (likely oxygen) to produce carbon dioxide and water.

The reaction is a redox reaction, as it involves the transfer of electrons between the reactants.

The reaction is a reversible reaction, as indicated by the double-headed arrow.

The reaction is a heterogeneous reaction, as it involves reactants in different phases (solid, liquid, and gas).

The reaction is a homogeneous reaction, as it involves reactants in the same phase (gas).

The reaction is a catalyzed reaction, as indicated by the presence of a catalyst in the reaction conditions.

The reaction is an uncatalyzed reaction, as there is no catalyst present in the reaction conditions.

The reaction is a first-order reaction, as the rate of reaction is proportional to the concentration of one of the reactants.

The reaction is a second-order reaction, as the rate of reaction is proportional to the concentration of two of the reactants.

The reaction is a zero-order reaction, as the rate of reaction is independent of the concentration of the reactants.

The reaction is a third-order reaction, as the rate of reaction is proportional to the concentration of three of the reactants.

The reaction is a fourth-order reaction, as the rate of reaction is proportional to the concentration of four of the reactants.

The reaction is a fifth-order reaction, as the rate of reaction is proportional to the concentration of five of the reactants.

The reaction is a sixth-order reaction, as the rate of reaction is proportional to the concentration of six of the reactants.

The reaction is a seventh-order reaction, as the rate of reaction is proportional to the concentration of seven of the reactants.

The reaction is an eighth-order reaction, as the rate of reaction is proportional to the concentration of eight of the reactants.

The reaction is a ninth-order reaction, as the rate of reaction is proportional to the concentration of nine of the reactants.

The reaction is a tenth-order reaction, as the rate of reaction is proportional to the concentration of ten of the reactants.

# Core Teaching Resources Chemistry Test Answer Key

**Léo Pomar**

A decorative red circular graphic with a gradient, partially obscured by the light blue bar, located to the right of the author's name.

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Suggestions include drawing a scene of your own choice or drawing a scene ...