

___K ___T ___C ___A
 20 5 8 8

Assessment of Learning (Product)

Multiple Choice: Circle the most correct answer. (20K)

1. The element X has three naturally occurring isotopes. The isotopic masses (amu) and % abundances of the isotopes are given in the table below. The average atomic mass of the element is _____ amu.

Isotope	Abundance	Mass
^{55}X	19.61	52.62
^{56}X	53.91	56.29
^{58}X	26.48	58.21

- A) 33.33 B) 55.74 C) 56.11 D) 57.23 E) 56.29

2. The seven elements that occur as diatomic elements are:

- A) H_2 , N_2 , O_2 , He_2 , Ne_2 , Cl_2 , Br_2 B) H_2 , N_2 , O_2 , He_2 , Ne_2 , C_2 , Na_2
 C) H_2 , N_2 , O_2 , F_2 , Cl_2 , Br_2 , I_2 D) Fe_2 , Ru_2 , O_2 , He_2 , Ne_2 , C_2 , Br_2

3. What is the total number of electrons in a Mg^{2+} ion?

- A) 10 B) 14 C) 12 D) 24

4. Atoms of which element have the greatest tendency to gain electrons?

- A) bromine B) fluorine C) chlorine D) iodine

5. Use the periodic table to predict which element has the largest ionization energy.

- A) H B) Kr C) Fr D) Rn E) Rh (# 45)

6. As you move across a period, from left to right, the size of the atoms of each element:

- A) increases B) decreases C) stays the same D) none of the above

7. The least electronegative elements are the:

- A) Alkali metals B) Metalloids C) Halogens D) Noble gases

8. In which one of these parts of the periodic table are transition metals located?

- A) in groups 1 & 2 at the left side of the table
 B) in group 18 at the right side of the table
 C) in groups 3-12 across the middle of the table
 D) in the lanthanoid and actinoid series

9. An atom has an atomic number of 16 and a mass number of 32. The number of electrons in the third orbit of the atom is:

- A) 2 B) 3 C) 5 D) 6 E) 8

Core Teaching Resources Chemistry The Periodic Table Test

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Resources in Education ,1998 **Teaching and Learning Online** Franklin S. Allaire,Jennifer E. Killham,2023-01-01

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