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mass C_6H_6 6 12 6 1 g mol 1 78 g mol 1 hence C_6H_6 s number of moles $\frac{30.78 \text{ mol}}{0.3846 \text{ mol}}$ molar mass of carbon tet CCl_4 1 12 4 355 154 g mol 1

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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