

## CHARLES' LAW

Name \_\_\_\_\_

Charles' Law states that the volume of a gas varies directly with the Kelvin temperature, assuming that pressure is constant. We use the following formulas:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad \text{or} \quad V_1 \times T_2 = V_2 \times T_1$$

$$K = ^\circ C + 273$$

Solve the following problems assuming a constant pressure.

1. A sample of nitrogen occupies a volume of <sup>25 L</sup>~~250 mL~~ at 25° C. What volume will it occupy at 95° C?  
\_\_\_\_\_
2. Oxygen gas is at a temperature of 40° C when it occupies a volume of 2.3 liters. To what temperature should it be raised to occupy a volume of 6.5 liters?  
\_\_\_\_\_
3. Hydrogen gas was cooled from 150° C to 50° C. Its new volume is <sup>7.5 L</sup>~~75 mL~~. What was its original volume?  
\_\_\_\_\_
4. Chlorine gas occupies a volume of <sup>2.5 L</sup>~~25 mL~~ at 300 K. What volume will it occupy at 600 K?  
\_\_\_\_\_
5. A sample of neon gas at 50° C and a volume of 2.5 liters is cooled to 25° C. What is the new volume?  
\_\_\_\_\_
6. Fluorine gas at 300 K occupies a volume of <sup>50 L</sup>~~500 mL~~. To what temperature should it be lowered to bring the volume to <sup>30 L</sup>~~300 mL~~?  
\_\_\_\_\_
7. Helium occupies a volume of 3.8 liters at -45° C. What volume will it occupy at 45° C?  
\_\_\_\_\_
8. A sample of argon gas is cooled and its volume went from <sup>30 L</sup>~~300 mL~~ to <sup>25 L</sup>~~250 mL~~. If its final temperature was -55° C, what was its original temperature?  
\_\_\_\_\_

# Boyles Law Chemistry If8766 Answers With Work

**AN Whitehead**



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