



Name:

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Student Exploration: Bohr Model: Introduction

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: absorption spectrum, Bohr model, electron volt, energy level, laser, orbit, photon

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

When light passes through a gas, certain wavelengths of the light are absorbed. The result is a unique **absorption spectrum**. Two examples are shown below.



1. What colors of light are absorbed by hydrogen gas?

The colors encircled by the vertical black lines are those that hydrogen gas has absorbed.

2. What colors of light are absorbed by helium gas?

The colors surrounded by the vertical black lines are those that helium gas has absorbed.

Gizmo Warm-up

In 1913, Niels Bohr proposed that the unique spectral lines created by different elements were related to the way electrons were arranged around the nucleus. The *Bohr Model: Introduction* Gizmo explores this connection.

The **laser** shown in the Gizmo can emit **photons**, or particles of light, at a variety of wavelengths. The energy of a photon, measured in **electron volts** (eV), is inversely proportional to its wavelength. Photons that pass through the gas are detected by the photon detector at right.



1. With the **Energy (eV)** set to 1 eV, click **Fire**. Did the photon go straight through the gas in the tube, or was it absorbed by the gas?
2. Set the **Energy (eV)** to 4 eV, and click **Fire**. What happened this time?

The photon was absorbed by the gas.

Bohr Model Gizmo Answer Key

Jin-Ying Zhang

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Bohr Model Gizmo Answer Key:

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